

Reactions of Amines and Epoxides

Preparation of Amines

Amination of alcohols: Polyalkylation is a serious problem for tertiary hydroxyl group



A summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



Bond Types, Strength, Length, Polarity and Dipole Moment

Two types of covalent bond

There are two types of covalent bond: ionic bond and covalent bond

- Ionic bonds are formed by the transfer of electrons from one element to another (ionization potential) and electron affinity (electron affinity).



- Covalent bonds are formed by sharing of electrons between two atoms.



The Relative Electronegativity of an



Bond and Bonding in a molecule

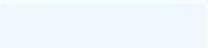


Bonding in a molecule



There are two types of covalent bond

Definition of covalent bond



Sigma and pi bonds

Covalent bonds are classified into sigma and pi bonds

- Sigma bonds are formed by head-on overlap of orbitals.
- pi bonds are formed by side-on overlap of orbitals.
- One single bond is a sigma bond.
- One multiple bond consists of one sigma and pi bonds.

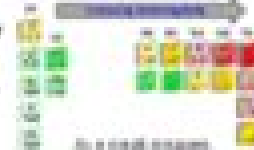


pi bonds are stronger than sigma bonds

Polar and Nonpolar Covalent Bonds

Depending on the difference in the electronegativity

covalent bonds can be polar and nonpolar.



- If the difference in electronegativity is less than 0.5, the electrons are shared equally (nonpolar bond).
- If the difference in electronegativity is between 0.5 and 1.7, the electrons are shared unequally (polar bond).
- If the difference in electronegativity is greater than 1.7, the electrons are not shared (ionic bond).

- A difference of 1.7 or higher can be high for the electrons are not shared (ionic bond).

- A difference of 1.7 or higher can be high for the electrons are not shared (ionic bond).

Dipole Moment

The polarity of a bond is described by the dipole moment (μ).



- The dipole moment is a vector pointing from the less electronegative to the more electronegative atom.
- The vector dipole moment is a vector pointing from the less electronegative to the more electronegative atom.
- μ (dipole moment) indicates an electron shift.
- δ (dipole moment) indicates an electron shift.

Molecular Dipole

To determine the molecular polarity, the vector sum of the dipole moment

The polarity of the molecule is a vector sum of the dipole moment. The polarity of the molecule is a vector sum of the dipole moment.



Vector sum of the dipole moment is the molecular dipole moment

Bond Length and Bond Strength

The bond length is the distance between the nuclei of the two atoms.

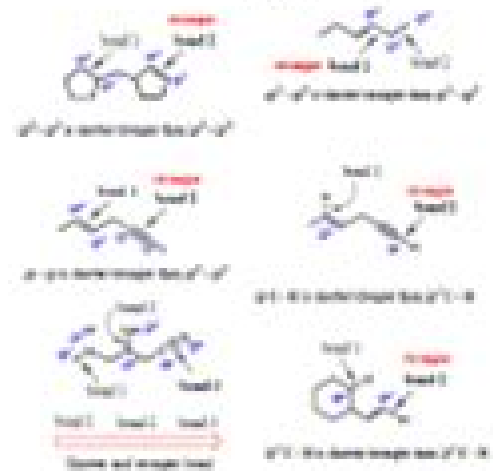


The bond length is the distance between the nuclei of the two atoms.

Average values of C-C and C-H bonds depending on hybridization			
C-C Bond	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Hybridization	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Length	1.54 Å	1.34 Å	1.20 Å
Strength	33 kcal/mol	47 kcal/mol	58 kcal/mol
Hybridization	sp^3-sp^3	sp^2-sp^2	$sp-sp$
C-H Bond	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Length	1.09 Å	1.08 Å	1.06 Å
Strength	98 kcal/mol	110 kcal/mol	120 kcal/mol

Higher s character means stronger bond

Example



Organic Reactions Study Guide

Paul A. Bartlett, Judith G. Koch



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Study Guide and Solutions Manual for Organic Chemistry Susan McMurry, 1992 John McMurry s best selling text presents organic chemistry in a new edition that is up to date beautifully written visually striking and pedagogically sound Described by many of its users as an eminently teachable text McMurry sets the standard in the field The writing style has received almost universal acclaim from its users McMurry introduces new concepts only as

needed and immediately illustrates them with concrete examples And wherever possible he ties material together with brief reviews overviews and reaction summaries The result is a text that helps students mentally organize the material a text that helps them understand concepts not just memorize facts and a text that helps them make sense of the voluminous amount of material they encounter in the study of organic chemistry McMurry uses a simple but important polar reaction the addition of HBr to an alkene as the lead off reaction to illustrate the general principles of organic reactions Users of former editions found this an excellent choice because of its relative simplicity no prior knowledge of chirality or kinetics is required and its importance as a polar reaction on a common functional group that offers students the key to understanding hundreds of thousands of ionic reactions By selecting this particular model McMurry is able to offer an unusually early presentation of organic reactions

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