

Reactions of Amines and Epoxides

Preparation of Amines

Amines are colorless, basic, and have a strong, fishy odor. They are prepared by the reduction of nitriles.



A summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



Epoxide Reactions and Mechanisms

Bond Types, Strength, Length, Polarity and Dipole Moment

Free-Radical Reactions

There are two types of bonding: ionic bonding and covalent bonding.

- Ionic bonds are formed by the transfer of electrons from one element to another (ionization potential) and electron affinity (electron).



- Covalent bonds are formed by sharing of electrons between two atoms.



The Relative Strength of Ions



Bond and Reaction Rates



Bonding in Molecules



There are two types of bonding

Definition of ionic and covalent bonds



Sigma and pi bonds

Covalent bonds are classified as sigma and pi bonds.

- Sigma bonds are formed by head-on overlap of orbitals.
- pi bonds are formed by side-on overlap of orbitals.
- A single bond is a sigma bond.
- A double bond consists of one sigma and one pi bond.



pi bonds are stronger than sigma bonds.

Polar and Nonpolar Covalent Bonds

Depending on the difference in the electronegativity values, covalent bonds can be polar and nonpolar.



• If the difference in electronegativity is less than 0.5, the electrons are shared equally (nonpolar bond). For example, the C-C and C-H bonds.

• If the difference in electronegativity is between 0.5 and 1.7, the electrons are shared unequally (polar bond). For example, the C-O and C-N bonds.

• A difference of 1.7 or higher is ionic, and the electrons are not shared (ionic bond).

Dipole Moment

The polarity of a bond is described by the dipole moment (μ).

It is a vector quantity.



The direction of the dipole moment is from the positive end to the negative end.

- The dipole moment is a vector pointing from the positive end to the negative end.
- The vector dipole moment is a vector pointing from the positive end to the negative end.
- $\mu = q \times r$ (where q is the charge and r is the distance between the charges).
- μ (dipole moment) indicates the direction of the bond.

Molecular Dipole

To determine the molecular dipole, the vector sum of the dipole moments is used.

The polarity of the molecule is a vector sum of the dipole moments of the bonds. A good example is carbon dioxide.



Carbon dioxide is a nonpolar molecule.

Bond Length and Bond Strength

The bond length is the distance between the nuclei of the two atoms.

| Bond | H-H | H-Cl | H-Br | H-I |
|----------|------------|------------|------------|------------|
| Length | 109 pm | 127 pm | 141 pm | 154 pm |
| Strength | 436 kJ/mol | 431 kJ/mol | 366 kJ/mol | 298 kJ/mol |

Double bonds are stronger than single bonds.

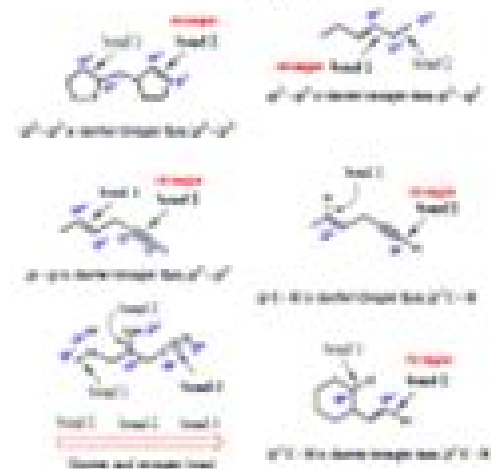


The bond length is the distance between the nuclei of the two atoms.

| Average values of C-C and C-H bonds depending on hybridization | | | |
|--|-------------|-------------|------------|
| C-C Bond | sp^3-sp^3 | sp^2-sp^2 | $sp-sp$ |
| Hybridization | sp^3-sp^3 | sp^2-sp^2 | $sp-sp$ |
| Length | 154 pm | 134 pm | 121 pm |
| Strength | 347 kJ/mol | 461 kJ/mol | 591 kJ/mol |
| Hybridization | sp^3-sp^2 | sp^2-sp | $sp-sp$ |
| Length | 147 pm | 133 pm | 121 pm |
| Strength | 356 kJ/mol | 475 kJ/mol | 591 kJ/mol |

Higher s character means stronger bond

Examples



Organic Chemistry Reactions Study Guide

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understood Each chapter ends with a list of frequently asked questions and answers followed by additional practice problems Answers are included in the Appendix

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This case clearly demonstrates the importance of synthesis and the use of organic chemistry. It's just one of the many examples used in the text that will spark the interest of students and get them involved in the study of organic chemistry.

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