

Sample Question Paper-4

CHEMISTRY (862)

Class-12

SOLVED

Time Allowed : 3 Hours

Maximum Marks : 70

This paper is divided into four sections – A, B, C and D.

Answer all questions.

Section A consists of one question having sub-parts of one mark each.

Section B consists of ten questions of two marks each.

Section C consists of seven questions of three marks each, and

Section D consists of three questions of five marks each.

Internal choices have been provided in one question each in Section B, Section C and Section D.

All working, including rough work, should be done on the same sheet as, and adjacent to the rest of the answer.

The intended marks for questions or parts of questions are given in brackets [].

Balanced equations must be given wherever possible and diagrams where they are helpful.

When solving numerical problems, all essential working must be shown.

In working out problems, use the following data:

Gas constant $R = 1.987 \text{ cal deg}^{-1} \text{ mol}^{-1} = 8.314 \text{ JK}^{-1} \text{ mol}^{-1}$

$= 0.0821 \text{ dm}^3 \text{ atm K}^{-1} \text{ mol}^{-1}$

$1.1 \text{ atm} = 1 \text{ dm}^3 \text{ atm} = 101.3 \text{ J}$, $1 \text{ Faraday} = 96500 \text{ coulombs}$.

Avogadro's number $= 6.022 \times 10^{23}$.

Section - A

(14 Marks)

Question 1

- (A) Fill in the blanks by choosing the appropriate word(s) from those given in the brackets. [4 × 1]
- (i) Ethyl isocyanide, on hydrolysis with dilute sulphuric acid gives and (formic acid/ acetic acid/ethylamine)
- (ii) In coordination complexes, the central metal atom or ion behaves as and the ligands behave as (Coordination compounds/Lewis base/Lewis acid/double salts)
- (iii) Halosilanes undergoes reaction. [A]
- (iv) Aromatic ether is prepared by heating with (Aliphatic halide/ alkyl halide/aromatic oxide/aryl halide)
- (B) Select and write the correct alternative from the choices given below: [7 × 1]
- (i) For a spontaneous reaction ΔG° and ΔE° cell will be respectively:
(A) -ve and +ve (B) +ve and -ve (C) +ve and +ve (D) -ve and -ve
- (ii) Which of the following pairs represents linkage isomers? [A]
- (A) $[\text{Pd}(\text{PPh}_3)_2(\text{NCS})_2]$ and $[\text{Pd}(\text{PPh}_3)_2(\text{SCN})_2]$
(B) $[\text{Co}(\text{NH}_3)_5\text{NO}_2]\text{SO}_4$ and $[\text{Co}(\text{NH}_3)_5\text{SO}_4]\text{NO}_2$
(C) $[\text{PtCl}_2(\text{NH}_3)_4]\text{Br}_2$ and $[\text{Pt}(\text{Br})_2(\text{NH}_3)_4]\text{Cl}_2$
(D) $[\text{Co}(\text{NH}_3)_4][\text{PtCl}_4]$ and $[\text{Pd}(\text{PPh}_3)_2(\text{SCN})_2]$
- (iii) The actinoids show more oxidation states in general than the lanthanoids as:
(A) the 5f orbitals are more buried than 4f orbitals.
(B) there is similarity between 4f and 5f orbitals in their angular part of the wave function.

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