

Table K
Selected Polyatomic Ions

Formula	Name	Formula	Name
ClO_4^-	perchlorate	ClO_2^-	chlorite
ClO_3^-	chlorate	ClO_2^+	dichlorine oxide
NH_4^+	ammonium	MnO_4^-	permanganate
$\text{C}_2\text{H}_3\text{O}_2^-$ CH_3COO^-	acetate	NO_2^-	nitrite
CN^-	cyanide	NO_3^-	nitrate
CO_3^{2-}	carbonate	O_2^{2-}	peroxide
HCO_3^-	hydrogen carbonate	OH^-	hydroxide
$\text{C}_2\text{O}_4^{2-}$	oxalate	PO_4^{3-}	phosphate
CrO_4^{2-}	chromate	SCN^-	thiocyanate
CrO_2^-	chromite	SO_3^{2-}	sulfite
CrO_4^{2-}	chromate	SO_3^{2-}	sulfite
$\text{Cr}_2\text{O}_7^{2-}$	dichromate	HSO_4^-	hydrogen sulfate
Fe_3O_4	pyrrole	$\text{S}_2\text{O}_8^{2-}$	persulfate

Table L
Solubility Guidelines for Aqueous Solutions

Basic Most Forms Soluble Compounds	Exceptions	Basic Most Forms Soluble Compounds*	Exceptions
Group 1 ions (Li^+ , Na^+ , etc.)		acetate (CH_3CO_2^-)	when combined with Group 1 ions or ammonium (NH_4^+)
ammonium (NH_4^+)		chromate (CrO_4^{2-})	when combined with Group 1 ions, Ca^{2+} , Sr^{2+} , or ammonium (NH_4^+)
nitrate (NO_3^-)		phosphate (PO_4^{3-})	when combined with Group 1 ions or ammonium (NH_4^+)
hydrogen carbonate (HCO_3^-)		oxide (O^{2-})	when combined with Group 1 ions or ammonium (NH_4^+)
halides (F^- , Cl^-)		sulfide (S^{2-})	when combined with Group 1 ions, Ca^{2+} , Sr^{2+} , Ba^{2+} , or ammonium (NH_4^+)
halides (Br^- , I^-)	when combined with Ag^+ , Hg_2^{2+} , or Pb^{2+}		
sulfates (SO_4^{2-})	when combined with Ag^+ , Ca^{2+} , Sr^{2+} , Ba^{2+} , or Pb^{2+}		

*Ammonium is always very low solubility in H_2O

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