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Mercury 50 Hp 2 Stroke Owners Manual

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Mercury 50 Hp 2 Stroke Owners Manual Introduction

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mole H_2S x 18g H_2O i mole H_2O 149 4g H_2O is your theoretical yield so percent yield $\frac{137 \text{ g } \text{H}_2\text{O}}{149 \text{ g } \text{H}_2\text{O}} \times 100 = 91.77\%$ problems 6 8 are similar to problem 5

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web a write a balanced equation for the reaction $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$ 4 Al + 3 $\text{O}_2 \rightarrow 2 \text{Al}_2\text{O}_3$ b determine the theoretical yield of Al_2O_3 $\frac{160 \text{ g Al}}{27 \text{ g/mol}} \times \frac{1 \text{ mol Al}}{2 \text{ mol Al}_2\text{O}_3} \times 101.96 \text{ g } \text{Al}_2\text{O}_3 = 299.8 \text{ g}$ $\frac{160 \text{ g Al}}{27 \text{ g/mol}} \times \frac{1 \text{ mol Al}}{2 \text{ mol Al}_2\text{O}_3} \times 101.96 \text{ g } \text{Al}_2\text{O}_3 = 299.8 \text{ g}$ $\frac{160 \text{ g Al}}{27 \text{ g/mol}} \times \frac{1 \text{ mol Al}}{2 \text{ mol Al}_2\text{O}_3} \times 101.96 \text{ g } \text{Al}_2\text{O}_3 = 299.8 \text{ g}$ c determine the percent yield $\frac{260 \text{ g}}{299.8 \text{ g}} \times 100 = 86.7\%$

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