

Unit 7: Reduction, Oxidation and ElectrochemistryChapter 17: Electrochemistry4.9: Oxidation-Reduction Reactions

Reduction-Oxidation Reactions (Redox Rxn): - chemical reactions where there is a transfer of electron(s).

Oxidation States (Oxidation Number): - a number that is arbitrary assigned to an atom in an element, molecule, and polyatomic ions to account for the number of electrons in *redox* reaction.

Rules for Assigning Oxidation Numbers:

1. Natural Elements (include diatomic and polyatomics) have an Oxidation Number of 0.

**Examples:** Na<sub>(s)</sub>, O<sub>2(g)</sub>, O<sub>3(g)</sub>, H<sub>2(g)</sub>, F<sub>2(g)</sub>, P<sub>4(s)</sub>, and Hg<sub>(l)</sub> all have an oxidation number of 0.

2. Single Atomic Ions have an Oxidation Number Equals to its Charge.

**Example:** K<sup>+</sup> has an oxidation number of +1.

3. Oxygen in Binary Compound and Polyatomic Ions has an Oxidation Number of -2.

**Examples:** CO<sub>2</sub>, CO, SO<sub>3</sub>, SO<sub>2</sub>, CO<sub>3</sub><sup>2-</sup>, SO<sub>4</sub><sup>2-</sup> and H<sub>2</sub>O all have an oxidation number of -2 for oxygen.

**An exception occurs in peroxides (compound containing O<sub>2</sub><sup>2-</sup>) where the oxygen has an oxidation number of -1.**

4. Hydrogen in Binary Compounds and Polyatomic Ions has an Oxidation Number of +1.

**Examples:** H<sub>2</sub>O, HCl, HBr, CH<sub>4</sub>, NH<sub>4</sub><sup>+</sup> and H<sub>2</sub>S all has an oxidation number of + 1 for hydrogen.

5. In Binary Compounds, the Atom that has the most attraction to electrons (HIGHER Electronegativity) is assigned the NEGATIVE Oxidation Number. This Negative Oxidation Number is the SAME as its Ionic Charge.

**Examples:** HF: F is more electronegative. Thus, F has an oxidation number of -1.

NH<sub>3</sub>: N is more electronegative. Thus, N has an oxidation number of -3.

H<sub>2</sub>S: S is more electronegative. Thus, S has an oxidation number of -2.

6. The Sum of all Oxidation Numbers is 0 for all Electrically Neutral Compounds.

**Examples:** CO: Oxidation numbers: O = -2; C = +2 [(+2) + (-2) = 0]

SO<sub>2</sub>: Oxidation numbers: O = -2; S = +4 [(+4) + 2(-2) = 0]

7. The Sum of all Oxidation Numbers is Equal to the Overall Charge for all Polyatomic Ions.

**Examples:** CO<sub>3</sub><sup>2-</sup>: Oxidation numbers: O = -2; C = +4 [(+4) + 3(-2) = -2]

SO<sub>4</sub><sup>2-</sup>: Oxidation numbers: O = -2; S = +6 [(+6) + 4(-2) = -2]

**Note:** Same Atom from different chemical species might NOT have the Same Oxidation Number.

We write the charges of ionic species as *n*<sup>+</sup> or *n*<sup>-</sup>. Oxidation numbers are written as +*n* or -*n*.

# Oxidation And Reduction Study Guide Answers

**Arshad Iqbal**



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## **Table of Contents Oxidation And Reduction Study Guide Answers**

1. Understanding the eBook Oxidation And Reduction Study Guide Answers
  - The Rise of Digital Reading Oxidation And Reduction Study Guide Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Oxidation And Reduction Study Guide Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Oxidation And Reduction Study Guide Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Oxidation And Reduction Study Guide Answers
  - Personalized Recommendations
  - Oxidation And Reduction Study Guide Answers User Reviews and Ratings
  - Oxidation And Reduction Study Guide Answers and Bestseller Lists
5. Accessing Oxidation And Reduction Study Guide Answers Free and Paid eBooks
  - Oxidation And Reduction Study Guide Answers Public Domain eBooks
  - Oxidation And Reduction Study Guide Answers eBook Subscription Services
  - Oxidation And Reduction Study Guide Answers Budget-Friendly Options
6. Navigating Oxidation And Reduction Study Guide Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Oxidation And Reduction Study Guide Answers Compatibility with Devices
  - Oxidation And Reduction Study Guide Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Oxidation And Reduction Study Guide Answers
  - Highlighting and Note-Taking Oxidation And Reduction Study Guide Answers
  - Interactive Elements Oxidation And Reduction Study Guide Answers
8. Staying Engaged with Oxidation And Reduction Study Guide Answers

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Oxidation And Reduction Study Guide Answers
- 9. Balancing eBooks and Physical Books Oxidation And Reduction Study Guide Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Oxidation And Reduction Study Guide Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Oxidation And Reduction Study Guide Answers
  - Setting Reading Goals Oxidation And Reduction Study Guide Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Oxidation And Reduction Study Guide Answers
  - Fact-Checking eBook Content of Oxidation And Reduction Study Guide Answers
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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