

Classify each statement about effective nuclear charge,  $Z_{\text{eff}}$ , as true or false.

True

Electrons in an  $s$  orbital are more effective than those in other orbitals at shielding other electrons from the nuclear charge.

In a Be atom, a  $1s$  electron has a greater  $Z_{\text{eff}}$  than a  $2s$  electron.

Effective nuclear charge decreases from left to right across a period on the periodic table.

False

Across a period, as  $Z_{\text{eff}}$  increases, atomic size decreases.

Effective nuclear charge does not depend on the number of electrons present in an atom.

A  $1s$  electron in a B atom has a smaller  $Z_{\text{eff}}$  than a  $1s$  electron in a Be atom.

Answer Bank

# Nuclear Chemistry Answers True Or False

**S. Chand Experts**



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