

Electrons particles in the nucleus of atoms.

1. negative
2. small

Atomic Number number of protons in the nucleus of an atom.

Electron shells all of _____ (1) = all of _____ (2)

Isotopes atoms of an element that have different numbers of neutrons.

Isotopes of Hydrogen

Hydrogen-1

1 proton and one electron

Hydrogen-2

1 proton and one electron

Hydrogen-3

1 proton and one electron

Mass Number total number of protons and neutrons in an atom.

Example: Carbon-12 6 protons 6 neutrons

Isotope	Protons	Neutrons	Mass Number
Hydrogen-1	1	0	1
Hydrogen-2	1	1	2
Hydrogen-3	1	2	3

Atomic Mass average of the masses of all the isotopes of an element.

Continued on: Extra Testing Study

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