Final Exam Ver. 1 (90 Multiple Choice)

Part A. Multiplie Choice: (90 marks) Edentify the letter of the choice that best completes the statement or asswers the question. Identify the compound that is likely to have the highest boding point? ALCOHOL: The ACCHIOCHT C. CHACH. 4. CRECHECHECHEOR W. CHACHACHACHACHACHA What forms when an alcohol is deliveraged? ALL AND STREET the age ascists the same additionable All of the automorphisms section as has been been as 3. What results when a secondary alcohol is exidized? A. Barrisoner All Addition the attlicence er. seminar Co. all delicities of the 4. Both copper (atomic number 2%) and chromium (atomic number 24) appear to break the pattern in the order of filling the 3d and 4s orbitals. This change in pattern is expressed by a, an increase in the number of electrons in both the 3d and 4s orbitals. b, a reduction in the number of electrons in both the 3d and 4s orbitals. c. a reduction in the number of electrons in the 3d orbital and an increase in the 4s orbital. d. a reduction in the number of electrons in the 4s orbital and an increase in the 3d orbital. 5. An electron has the following set of quantum numbers: $n = 3, l = 1, m_1 = 1, m_2 = \frac{n-2}{2}$. In which orbital is this electron found? de Des 46. 34 Boy Beet 400 400 Co. 2540 6. Which of the following groups of ions and atoms is comprised of species having exactly the same ground state electron configuration? A. All O' No. Cil Bury Call, Till L. Cill., Sci. c. D. He, Li, Be'll el. Car. Thr. Ch. St. e. No. Ar. Kr. No. Which condensed electron configuration represents the element. Bohrium (Z=1979) a. [Re[7s/56"4d" the different vision to the con- I PARTS 447 Set AL TRACTORALITY (Rai/25/54*66*) 8. Which type of hybridization is most likely to have occurred for the central atom in PCL? ALC: NOTE: eth. sept." the new left. 46 Sept. Mary Sept. Which electron configuration represents an alkali metal? a. 15° 25° 25° 35° 36°

3.6° 26° 20° 36° 36° 36°
 3.6° 20° 20° 36°

Night Final Exam Multiple Choice Version Answers

Robert L. Bratton

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