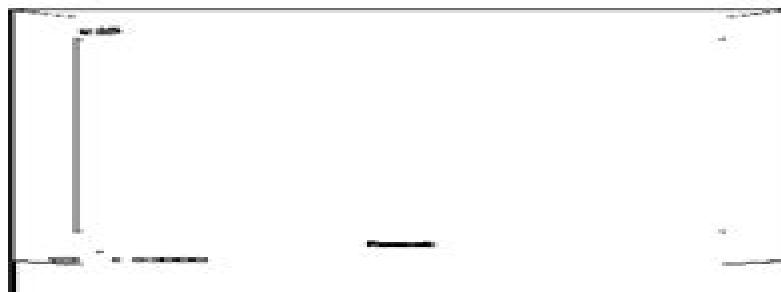


Service Manual

Plasma Television



TH-37PA60E
TH-37PD60B
TH-42PA60E
TH-42PD60B
 GP9DE Chassis

Specifications

Power Source	AC 220-240 V, 50 / 60 Hz
Power Consumption	
Average use	206 W (TH-37PA60E), 215 W (TH-37PD60B) 238 W (TH-42PA60E), 247 W (TH-42PD60B)
Standby condition	0.3 W (PA60E) 0.3 W (PD60B Without DVB recording), 20 W (PD60B With DVB recording)
Plasma Display panel	
Drive method	AC type
Aspect Ratio	16:9
Contrast Ratio	Max. 10000:1
Visible screen size (No. of pixels)	94 cmV (37-inch), 106 cmV (42-inch) 518 mm (W) × 461 mm (H) × 539 mm (diagonal) (37-inch) 620 mm (W) × 518 mm (H) × 1,056 mm (diagonal) (42-inch) 408,960 (852 (W) × 480 (H)) [2,556 × 480 dots]
Sound	
Speaker	12 cm × 6 cm × 2 pcs, 8 Ω
Audio Output	20 W (10 W + 10 W), 10% THD
Receiving Systems / Band name (PA60E)	PAL B, G, H, L, SECAM B, G, SECAM L / L' VHF E2 - E12 VHF A - H (ITALY) CATV S01 - S09 CATV S11 - S20 (U1 - U10) PAL D, K, SECAM D, K VHF R1 - R2 VHF R6 - R12 PAL S2/S60 VHF H1 - H2 (ITALY) UHF E21 - E59 CATV S1 - S30 (M1 - M10) CATV S21 - S41 (Hyperband) VHF R3 - R5 UHF E21 - E59

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Panasonic Th 42pa60e Th 42pd60b Service Repair Manual

David Kirk



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web k sp of calcium hydroxide 1 author j m mccormick last update ocober 13 2013 introduction an equilibrium constant k is related to Δg for a process through eqn 1 and as long as Δh and Δs are independent of temperature eqn 1 can be combined with eqn 2 to determine Δh and Δs from the temperature dependence of k

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web in today s experiment you will determine the solubility product ksp of calcium hydroxide ca oh 2 by measuring the concentration of ca oh 2 in a saturated solution calcium hydroxide is a sparingly soluble salt that dissolves according to the following reaction the solubility product expression for this reaction is

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web expert verified chem 182 determining the ksp of calcium hydroxide calcium hydroxide is an ionic solid that is sparingly soluble in water a saturated aqueous solution of ca oh is represented in equation form as shown below ca oh s \rightleftharpoons ca aq 2oh aq the solubility product expression describes in mathematical terms the equilibrium

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web use the concentration of calcium hydroxide to calculate the concentration of calcium ions and hydroxide ions in the solution ca oh 2 \rightleftharpoons ca 2 2oh 5 x 10⁻² m calcium ions 5 x 10⁻² m 2 0 m hydroxide ions calculate calcium hydroxide ksp using the concentrations of calcium ions and hydroxide ions

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web madison bowe elana keller april 15 2014 introduction the objective of this lab was to measure the ksp of a slightly soluble basic salt and how solubility changes with temperature ksp is another form of equilibrium expression that deals with

solubility the reaction that will take place is $\text{Ca(OH)}_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{OH}^-$

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web the chemical equation used was $\text{Ca(OH)}_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{OH}^-$ the solubility product was determined to be ksp Ca(OH)_2 from the titration which was made by dripping a concentrated solution of calcium hydroxide into hydrochloric acid the calcium hydroxide was filtered before starting the titration

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web jul 12 2023 because the concentration of a pure solid such as $\text{Ca}_3(\text{PO}_4)_2$ is a constant it does not appear explicitly in the equilibrium constant expression the equilibrium constant expression for the dissolution of calcium phosphate is therefore $K = [\text{Ca}^{2+}]^3 [\text{PO}_4^{3-}]^2$

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web therefore it was determined through error calculations and ease of use that the most reliable valid method to use to get the closest ksp value to the theoretical ksp for calcium was the determination through ph method 2 introduction calcium hydroxide is a strong base soluble in water

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web calculate the ksp for calcium hydroxide $\text{Ksp Ca}^{2+} \text{OH}^{-} = 2.0 \times 10^{-2}$ $\text{Ksp} = 5 \times 10^{-4}$ find the accepted value of the ksp for calcium hydroxide and compare it with your value discuss the discrepancy and suggest possible sources of experimental error accepted value for ksp for sodium hydroxide is 5×10^{-6}

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