



## Chapter 4 & 5: Modern Model of the Atom and Bonding

Supplemental Instruction  
Iowa State University

Leader:	Katie
Course:	Chem 163
Instructor:	Burnett/Kingston
Date:	September 28, 2009

### Review

- What is the noble gas configuration of the following atoms and ions?
  - Ar  $[\text{Ne}] 3s^2 3p^4$  or  $[\text{Ar}]$
  - $\text{Na}^+$   $[\text{Ne}] 2s^2 2p^6$  or  $[\text{Ne}]$
  - $\text{C}^{2-}$   $[\text{He}] 2s^2 2p^4$
  - $\text{O}^{2-}$   $[\text{He}] 2s^2 2p^6$  or  $[\text{Ne}]$
  - $\text{Ca}^{2+}$   $[\text{Ar}] 3s^2 3p^6$  or  $[\text{Ar}]$

### New Material

- Define the following terms and answer associated questions.
  - Heisenberg uncertainty principle:** Principle that states that the position and momentum of a subatomic particle cannot be known simultaneously.
    - orbital:** 3D volume in space, cloud like distribution of electrons.
      - What is its maximum capacity?  $2e^-$
    - covalent bond:** Chemical bond where the electrons are shared.
      - What does it represent? The covalent bond represents an energy minimum.
      - Is energy absorbed or released when a covalent bond is formed? Energy is released.
    - bond length:**
      - What causes bond length? The attraction of electrons to the nuclei and repulsion of 2 nuclei.
      - What units is bond length measured in?  $\text{\AA}$  = Angstrom =  $10^{-10} \text{ m}$  = 1m
  - How many orbitals are there in any s sublevel? Any p sublevel? Any d sublevel?
    - s = 1 orbital -  $2e^-$
    - p = 3 orbitals -  $6e^-$
    - d = 5 orbitals -  $10e^-$
  - Draw the orbitals for the  $2p^5$ . How many electrons are there in each orbital?



- Write the following atoms in Lewis Dot notation.
  - N  $\text{N}^\bullet$
  - H  $\text{H}^\bullet$
  - Xe  $\text{Xe}^\bullet$
  - Fe  $\text{Fe}^\bullet$

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