Solution Manual for Modern Semiconductor Devices for Integrated Circuits Chenming C. Hu

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Chapter 1

Visualization of the Silicon Crystal

- 1.1 (a) Please refer to Figure 1-2. The 8 corner atoms are shared by 8 unit cells and therefore contribute 1 atom. Similarly, the 6 face atoms are each shared by 2 unit cells and contribute 3 atoms. And, 4 atoms are located inside the unit cell. Hence, there are total 8 silicon atoms in each unit cell.
 - (b) The volume of the unit cell is

$$V_{\text{max cert}} = (5.43 \text{ A})^3 = (5.43 \times 10^{-8} \text{ cm})^3 = 1.60 \times 10^{-22} \text{ cm}^3$$
,

and one unit cell contains 8 silicon atoms. The atomic density of silicon is

$$N_{Si} = \frac{8 \ silicon \ atoms}{V_{min \ coll}} = 5.00 \times 10^{22} \ (silicon \ atoms) \ cm^{-3}$$
 .

Hence, there are 5.00×10²² silicon atoms in one cubic centimeter.

(c) In order to find the density of silicon, we need to calculate how heavy an individual silicon atom is

$$Mass_{1.52 \text{ seem}} = \frac{28.1 (g/mole)}{6.02 \times 10^{23} (atoms/mole)} = 4.67 \times 10^{-23} (g/atom).$$

Therefore, the density of silicon (psi) in g/cm3 is

$$\rho_{s_i} = N_{s_i} \times Mass_{1s_{i,mom}} = 2.33 \text{ g/cm}^3$$
.

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