

Mole Lab

Introduction to The Mole Concept



Introduction

Although technically not a laboratory experiment, this activity certainly helps to drive home the main idea behind the mole concept—that chemists can count out infinitesimally small particles by weighing.

Concepts

- Avogadro's number
- Chemical formulas
- Molar mass or molecular weight

Materials (for each lab station, two or three lab pairs can work at each station)

- Aluminum foil
- Ammonium sulfate, $(\text{NH}_4)_2\text{SO}_4$, 100 g
- Sodium chloride, NaCl , 500 g
- Water, distilled, 500 mL
- Balance, electronic, 0.01-g precision (if only triple beam balances are available, it is best if there is one for each lab pair)
- Beakers, 600-mL, or cut off 2-L soda bottles, 3 labeled
- Cups, blue, clear, and red plastic, 2–3 of each
- Pipets, Beral-type, 2–3
- Spoons, plastic, 4–6

Materials (for the instructor's station, preferably near the center of the room)

- Copper sheet, 12 pieces*
 - Balance, electronic, 0.01-g precision
 - Cups, blue, clear, and red plastic, 1 each
 - Scoring Standard Table
- *See the Preparation section below for sizes of copper pieces.*

Safety Precautions

Ammonium sulfate is slightly toxic by ingestion. Wear chemical splash goggles, chemical-resistant gloves, and a chemical-resistant apron. Wash hands thoroughly with soap and water before leaving the laboratory. Follow all laboratory safety guidelines. Please review current Material Safety Data Sheets for additional safety, handling, and disposal information.

Preparation

1. Label three 600-mL beakers “NaCl,” “water,” and “ammonium sulfate” and add to each the corresponding substance. *Note:* Do not write the chemical formula for ammonium sulfate on the beaker; students must determine this.
2. Place 2–3 spoons in each beaker containing a solid substance.
3. Place 2–3 pipets in the beaker of water.
4. Cut twelve pieces of copper sheet according to Table 1 and label with the corresponding numbers 1–12. *Note:* The copper pieces can have any weights you want them to. Using the weights given in the table allows you to use the Scoring Standard table as is. If you use other weights, just modify column C in the table accordingly, as shown in Figure 1 on page 2.

Cu Piece	Mass (g)
#1	4.22
#2	4.78
#3	4.98
#4	5.13
#5	5.44
#6	5.97
#7	6.09
#8	6.55
#9	6.79
#10	7.13
#11	7.65
#12	7.99

Mole Lab Student Guide

Terrence J. Bazzett, Bryan Kolb

A red circular graphic with a gradient, appearing as a partial circle or a stylized arrow pointing to the right, located on the right side of the slide.

Mole Lab Student Guide:

ChemDiscovery Student Guide Olga I. Agapova, 2002 **Chemistry, Student Study Guide** James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane CH_4 from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement *The Organic Chem Lab Survival Manual* James W. Zubrick, 2016-01-19 Written for the laboratory that accompanies the sophomore junior level courses in Organic Chemistry Zubrick provides students with a valuable guide to the basic techniques of the Organic Chemistry lab The book will help students understand and practice good lab safety It will also help students become familiar with basic instrumentation techniques and apparatus and help them master the latest techniques such as interpretation of infrared spectroscopy The guide is mostly macroscale in its orientation Student Study Guide and Solutions Manual to accompany General Organic and Biological Chemistry, 1e Kenneth W. Raymond, 2005-10-07 Finally readers have a shorter less intimidating introduction to general organic and biological chemistry Not only is Raymond's text concise it also takes an integrated approach to presenting important topics in a way that makes the material easier to understand In this approach similarities can be exploited and concepts reinforced The result is that readers see the strong connections that exist between these three branches of chemistry *Study Guide for Chemical Principles* Thomas Elliott Taylor, 1979 **The Complete Chemistry Study Guide and Note Cards and MCAT** Konstantinos Papadopoulos, 2012-07-06 **Barron's Science 360: A Complete Study Guide to Physics with Online Practice** Barron's Educational Series, Kenneth Rideout, 2021-09-07 Barron's Math 360 Physics is your complete go to guide for everything physics This comprehensive guide is an essential resource for High school and college courses Homeschooling Virtual Learning Learning pods Inside you'll find Comprehensive Content Review Begin your study with the basic building blocks of physics and build as you go Topics include motion forces electricity magnetism and introduction to nuclear physics and much more Effective Organization Topic organization and simple lesson formats break down the subject matter into manageable learning modules that help guide a successful study plan customized to your needs Clear Examples and Illustrations Easy to follow explanations hundreds of helpful illustrations and numerous step by step examples make this book ideal for self study and rapid learning Practice Exercises Each chapter ends with practice exercises designed to reinforce and extend key skills and concepts These checkup exercises along with the answers and solutions will help you assess your understanding and monitor your progress Access to

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