

Key.

b) 6.45 moles  $\text{Cr}_2\text{O}_3$   $\text{FW} = 152.0 \text{ g/mol}$  980 g

$$6.45 \text{ mol} \times \left( \frac{152.0 \text{ g}}{1 \text{ mol}} \right) = 980 \text{ g}$$

c) 0.144 mole  $\text{K}_2\text{S}$   $\text{FW} = 110.3 \text{ g/mol}$  15.9 g

$$0.144 \text{ mol} \times \left( \frac{110.3 \text{ g}}{1 \text{ mol}} \right) = 15.9 \text{ g}$$

d) 5.0 moles  $\text{C}_6\text{H}_{12}\text{O}_6$   $180.0 \text{ g/mol}$  900 g  $\approx 9.0 \times 10^2 \text{ g}$

$$5.0 \text{ mol} \times \left( \frac{180.0 \text{ g}}{1 \text{ mol}} \right) = 900 \text{ g}$$

e)  $5.10 \times 10^{-4}$  mole  $\text{NaCl}$   $\text{FW} = 58.5 \text{ g/mol}$   $0.0298 \text{ g} \approx 2.98 \times 10^{-2} \text{ g}$

$$5.10 \times 10^{-4} \text{ mol} \times \left( \frac{58.5 \text{ g}}{1 \text{ mol}} \right) = 0.0298 \text{ g}$$

Solve the following stoichiometry problems. Write a balanced equation for each.

a) How many moles of hydrogen gas are needed to react with 15.1 g chlorine gas, to produce hydrogen chloride gas? Equation:  $\text{H}_2(\text{g}) + \text{Cl}_2(\text{g}) \rightarrow 2 \text{HCl}(\text{g})$

$$15.1 \text{ g} \times \left( \frac{1 \text{ mol}}{71.0 \text{ g}} \right) = 0.213 \text{ mol } \text{Cl}_2$$

$$0.213 \text{ mol } \text{Cl}_2 \left( \frac{1 \text{ mol H}_2}{1 \text{ mol Cl}_2} \right) = 0.213 \text{ mol H}_2 \quad \underline{0.213 \text{ mol H}_2}$$

b) How many grams of oxygen gas are needed to react completely with 16.2 g of hydrogen gas, to produce water? Equation:  $2 \text{H}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2 \text{H}_2\text{O}(\text{l})$

$$16.2 \text{ g} \times \left( \frac{1 \text{ mol}}{2.0 \text{ g}} \right) = 8.10 \text{ mol H}_2$$

$$8.10 \text{ mol H}_2 \left( \frac{1 \text{ mol O}_2}{2 \text{ mol H}_2} \right) = 4.05 \text{ mol O}_2 \quad \underline{130 \text{ g O}_2}$$

$$4.05 \text{ mol O}_2 \left( \frac{32.0 \text{ g/mol}}{1 \text{ mol}} \right) = 130 \text{ g O}_2$$

c) How many moles of calcium oxide are produced when 36.5 g of calcium react completely with oxygen gas? Equation:  $2 \text{Ca}(\text{s}) + \text{O}_2(\text{g}) \rightarrow 2 \text{CaO}(\text{s})$

$$36.5 \text{ g} \left( \frac{1 \text{ mol}}{40.1 \text{ g}} \right) = 0.910 \text{ mol Ca}$$

$$0.910 \text{ mol Ca} \left( \frac{2 \text{ mol CaO}}{2 \text{ mol Ca}} \right) = 0.910 \text{ mol } \underline{0.910 \text{ mol CaO}}$$

d) How many grams of aluminum are required to produce 410 g of aluminum oxide, through reaction with oxygen gas? Equation:  $4 \text{Al} + 3 \text{O}_2 \rightarrow 2 \text{Al}_2\text{O}_3$

$$410 \text{ g} \left( \frac{1 \text{ mol}}{102.0 \text{ g}} \right) = 4.02 \text{ mol Al}_2\text{O}_3$$

$$4.02 \text{ mol Al}_2\text{O}_3 \left( \frac{4 \text{ mol Al}}{2 \text{ mol Al}_2\text{O}_3} \right) = 8.04 \text{ mol Al}$$

$$8.04 \text{ mol Al} \left( \frac{27.0 \text{ g}}{1 \text{ mol}} \right) = 217 \text{ g}$$

$$\underline{217 \text{ g} \approx 220 \text{ g (2sf)}}$$

e) How many grams of solid iodine are produced by the decomposition of 25.4 grams of hydrogen iodide gas? Equation:  $2 \text{HI} \rightarrow \text{H}_2 + \text{I}_2$

$$25.4 \text{ g} \left( \frac{1 \text{ mol}}{127.9 \text{ g}} \right) = 0.199 \text{ mol HI}$$

$$0.199 \text{ mol HI} \left( \frac{1 \text{ mol I}_2}{2 \text{ mol HI}} \right) = 0.0993 \text{ mol I}_2 \quad \underline{25.2 \text{ g I}_2}$$

$$0.0993 \text{ mol I}_2 \left( \frac{253.8 \text{ g}}{1 \text{ mol}} \right) = 25.2 \text{ g I}_2$$

# Mole Ratio Wkst Answers

**Saraswati Experts**

A red circular graphic with a gradient, appearing as a stylized arrow or a partial circle, located to the right of the 'Saraswati Experts' text.

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**Resources in Education** ,1994

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- Chapter 2: Essential Elements of Mole Ratio Wkst Answers
- Chapter 3: Mole Ratio Wkst Answers in Everyday Life
- Chapter 4: Mole Ratio Wkst Answers in Specific Contexts
- Chapter 5: Conclusion

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## **Table of Contents Mole Ratio Wkst Answers**

1. Understanding the eBook Mole Ratio Wkst Answers



- The Rise of Digital Reading Mole Ratio Wkst Answers
- Advantages of eBooks Over Traditional Books
- 2. Identifying Mole Ratio Wkst Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Mole Ratio Wkst Answers
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Mole Ratio Wkst Answers
  - Personalized Recommendations
  - Mole Ratio Wkst Answers User Reviews and Ratings
  - Mole Ratio Wkst Answers and Bestseller Lists
- 5. Accessing Mole Ratio Wkst Answers Free and Paid eBooks
  - Mole Ratio Wkst Answers Public Domain eBooks
  - Mole Ratio Wkst Answers eBook Subscription Services
  - Mole Ratio Wkst Answers Budget-Friendly Options
- 6. Navigating Mole Ratio Wkst Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Mole Ratio Wkst Answers Compatibility with Devices
  - Mole Ratio Wkst Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Mole Ratio Wkst Answers
  - Highlighting and Note-Taking Mole Ratio Wkst Answers
  - Interactive Elements Mole Ratio Wkst Answers
- 8. Staying Engaged with Mole Ratio Wkst Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Mole Ratio Wkst Answers

9. Balancing eBooks and Physical Books Mole Ratio Wkst Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Mole Ratio Wkst Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Mole Ratio Wkst Answers
  - Setting Reading Goals Mole Ratio Wkst Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Mole Ratio Wkst Answers
  - Fact-Checking eBook Content of Mole Ratio Wkst Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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