

Moles and Chemical Formulas

11

A. Finding the Simplest Formula

1. Mass of empty crucible + cover
2. Initial appearance of the magnesium
3. Mass of crucible + cover + magnesium
4. Mass of crucible + cover + oxide product

| | |
|--------|---|
| 22.842 | g |
| 23.483 | |
| 23.073 | g |
| 29.317 | g |

Calculations

5. Mass of magnesium
6. Mass of magnesium compound
7. Mass of oxygen in the product
8. Moles of Mg

| | |
|--------|---|
| 0.2305 | g |
| 6.445 | g |
| 6.245 | g |
| .0095 | g |

(Show calculations.)

$$2305 \text{ g/mg} \times \frac{1 \text{ mole mg}}{24.31 \text{ g/mg}} = .009481$$

9. Moles of O

(Show calculations.)

$$6.245 \text{ g} \times \frac{1 \text{ mole}}{16.00} = .3903$$

10. Which number of moles (Mg or O) is smaller Mole of Mg
moles of Mg (rounded)

Mole Lab Chemistry I Acc Answers

Peipei Pang



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