Overview Chemical Bonds

Directions: All of the statements below are false as written. In the space provided, write a term or phrase that makes the statement true when it is substituted for the underlined words.

- 1. The properties of a compound are the same as the properties of the elements that it contains.

 2. Superscript numbers in chemical formulas tell how many atoms of each element are found in a unit of compound.

 3. All the noble gases except helium have 18 electrons in
 - their outer energy level.

 4. A(n) chemical formula is the force that holds atoms
 - together in a compound.

 S. An ion is a(n) <u>neutral</u> particle that has either more or fewer electrons than protons.
 - 6. Oxidation numbers are written as subscripts.
 - A(n) covalent bond is the force of attraction between the opposite charges of the ions in an ionic compound.
 - 8. The charge on a compound is always positive.
 - Equal sharing of electrons in covalent bonds results in polar molecules.
 - 10. Only two identical atoms can share electrons unequally.
 - 11. A binary compound contains five different elements.
 - An oxidation number tells how many <u>protons</u> an atom must gain, lose, or share to become stable.
 - The oxidation number of the copper(II) ion is 3+.
 - 14. When writing chemical formulas, add superscripts so that the sum of the oxidation numbers equals ten.
 - 15. A polyatomic ion never has a positive or negative charge.
 - 16. The polyatomic ion SO₄2- is called the sulfide ion.

Overview Chemical Bonds Answer Key

K. S. Verma

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