



Name \_\_\_\_\_

## Photoelectric Effect Lab

### Background

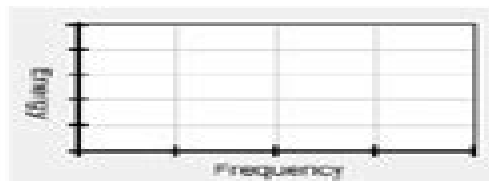
Physics was at its pinnacle of understanding the natural world as of 1900. Only three things had yet to be explained. Physics assumed once blackbody radiation, photoelectric effect and atomic spectra were explained that the study of physics would be complete. The assumption was wrong and led to a new branch of physics known today as quantum mechanics.

This lab will investigate the photoelectric effect. The photoelectric effect is the emission of electrons from a surface when illuminated with light of a certain frequency. The first insight to understanding this phenomenon was presented in 1900 by Max Planck. His formula,  $E = hf$ , related the energy of a photon to its frequency. Albert Einstein extended this idea of quantized photonic energy to a stream of photons (electromagnetic radiation) and explained the photoelectric effect.

### Part 1

Objective - What is the relationship between the energy of a photon and its frequency?

Qualitatively graph photonic energy versus frequency on the axes



to

### Part 2

Objective - What determines if electrons are ejected from the surface?

- Go to [http://phet.colorado.edu/simulations/sims.php?sim=Photoelectric\\_Effect](http://phet.colorado.edu/simulations/sims.php?sim=Photoelectric_Effect) and launch the simulation.
- Make the following adjustments to the simulation once it has launched.
  - Increase the intensity to 50%.
  - Check the box "electron energy vs light frequency".Once these adjustments have been made you should notice the ejection of electrons from the surface.
- Increase the wavelength of the light until electrons are no longer ejected. Record the wavelength in the table below and complete the calculations. 1 electron-volt (eV) =  $1.6 \times 10^{-19}$  J
- Repeat the above step for each of the metals under the pull down menu.

Metal	Wavelength (nm)	Frequency (Hz)	Energy (J)	Energy (eV)
Sodium				
Zinc				
Copper				
Platinum				
Calcium				

The

- minimum frequency of a photon that can eject an electron from a surface is called the threshold frequency,  $f_t$ . What is the threshold frequency,  $f_t$ , for each of the metals?

Metal	Threshold Frequency (Hz)
Sodium	
Zinc	
Copper	
Platinum	
Calcium	

### Part 3

# Modern Photoelectric Effect Lab Answers

**G Orfield**



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