

Molarity Practice problems

$$M = \frac{\text{mol}}{L}$$



Key

1. How many moles of CaCl_2 would be dissolved in 1.0 L of a 0.1 M solution of CaCl_2 ?

$$\text{mol} = M \times L = 0.1 \frac{\text{mol}}{L} \times 1 L = 0.1 \text{ mol } \text{CaCl}_2$$

2. What is the concentration of 25 mL of solution that has 0.05 moles of solute dissolved in it?

$$M = \frac{\text{mol}}{L} = \frac{0.05 \text{ mol}}{0.025 L} = 2 M$$

3. How many liters of water need to be added to 0.5 moles of CaCl_2 to make a 0.20 M solution of CaCl_2 ?

$$L = \frac{\text{mol}}{M} = \frac{0.5 \text{ mol}}{0.20 M} = 2.5 L$$

4. How many grams of potassium carbonate are needed to make 200 mL of a 2.5 M solution?

$$\text{mol} = M \times L \quad \text{K}_2\text{CO}_3$$

$$\text{mol} = 2.5 M \times .200 L = 0.50 \text{ mol}$$

$$\frac{0.50 \text{ mol } \text{K}_2\text{CO}_3}{1 \text{ mol}} \times 138.21 \text{ g} = 69.1 \text{ g } \text{K}_2\text{CO}_3$$

5. How many grams of Lithium Bromide would be needed to make 1.43 L of 4 M Lithium Bromide solution?

$$\text{mol} = M \times L \quad \text{LiBr}$$

$$= 4 M \times 1.43 L$$

$$= 5.72 \text{ mol } \text{LiBr}$$

$$\frac{5.72 \text{ mol } \text{LiBr}}{1 \text{ mol}} \times 86.84 \text{ g} = 496.7 \text{ g } \text{LiBr}$$

6. What is the concentration of a 450 mL solution that contains 200 grams of iron (II) chloride? FeCl_2

$$M = \frac{\text{mol}}{L}$$

$$\frac{200 \text{ g } \text{FeCl}_2}{126.75 \text{ g}} \times 1 \text{ mol} = 1.578 \text{ mol } \text{FeCl}_2$$

$$\frac{1.578 \text{ mol}}{.450 L} = 3.5 M$$

7. How many grams of ammonium sulfate are needed to make a 0.25 L solution at a concentration of 6 M?

$$\text{mol} = M \times L \quad (\text{NH}_4)_2\text{SO}_4$$

$$= 6 M \times .25 L$$

$$= 1.5 \text{ mol } (\text{NH}_4)_2\text{SO}_4$$

$$\frac{1.5 \text{ mol } (\text{NH}_4)_2\text{SO}_4}{1 \text{ mol}} \times 132.138 \text{ g} = 198.2 \text{ g } (\text{NH}_4)_2\text{SO}_4$$

8. What is the concentration of a solution that has a volume of 2.5 L and contains 660 grams of calcium phosphate?

$$M = \frac{\text{mol}}{L}$$

$$\frac{660 \text{ g } \text{Ca}_3(\text{PO}_4)_2}{310.18 \text{ g}} \times 1 \text{ mol} = 2.127 \text{ mol}$$

$$\frac{2.127 \text{ mol}}{2.5 L} = 0.85 M$$

9. How many grams of copper (II) fluoride are needed to make 6.7 liters of a 1.2 M solution?

$$\text{mol} = M \times L \quad \text{CuF}_2$$

$$= 1.2 M \times 6.7 L$$

$$= 8.04 \text{ mol } \text{CuF}_2$$

$$\frac{8.04 \text{ mol } \text{CuF}_2}{1 \text{ mol}} \times 101.55 \text{ g} = 816.5 \text{ g } \text{CuF}_2$$

Molarity Practice Problems Answer Key With Work

James E. Brady, John R. Holum



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Chemistry James E. Brady, John R. Holum, 1993-01-21 Offers accurate lucid and interesting explanations of basic concepts and facts of chemistry while helping students develop skills in analytical thinking and problem solving Students are taught in a variety of ways to think of skills as tools that can be used to solve complex problems Several aids are included to help focus and inspire student interest frequent reference to common chemicals in commercial products numerous photographs of reactions in chapter practice exercises following worked examples

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An Introduction to Chemistry Michael Mosher, Paul Kelter, 2023-03-18 This textbook is written to thoroughly cover the topic of introductory chemistry in detail with specific

references to examples of topics in common or everyday life It provides a major overview of topics typically found in first year chemistry courses in the USA The textbook is written in a conversational question based format with a well defined problem solving strategy and presented in a way to encourage readers to think like a chemist and to think outside of the box Numerous examples are presented in every chapter to aid students and provide helpful self learning tools The topics are arranged throughout the textbook in a traditional approach to the subject with the primary audience being undergraduate students and advanced high school students of chemistry

Mathematics for the Clinical Laboratory Lorraine J. Doucette,1997 This comprehensive practical text shows you how to perform the clinical calculations used in each area of the laboratory to ensure accurate results for patients

General Chemistry James P. Birk,Donald M. Baird,1994-03
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