VSEPR Theory (Molecular Shapes)

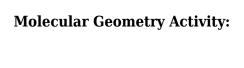
A = the central atom, X = an atom bonded to A, E = a lone pair on A

Note: There are lone pairs on X or other atoms, but we don't care. We are interested in only the electron densities or domains around atom A.

Total Domains	Generic Formula	Picture	Bonded Atoms	Lone Pairs	Molecular Shape	Electron Geometry	Example	Hybridi -zation	Bond Angles
	AX	A		0	Lirear	Linear	H ₂	•	180
2	AX ₂	X-A-X	2	0	Linear	Linear	CO ₃	sφ	180
	AXE	O A-X	1	- 1	Linear	Linear	CN'		
3	AX ₁	1	3	0	Trigonal planer	Trigonal planar	Albr ₃	se ²	120
	AX ₃ E	× ×	2	1	Bert	Trigonal planar	SnCl ₂		
	AHE,	-6		2	Linear	Trigonal planar	- Gi		
	HL.	1		8	Teterane	Totales	801,	ω'	100.5
	NJ.	2	3	1	Triposi pyramit	Yearen	Pis,		
	AKE,	1	1	2	Bell	Television	Salt-		
	AND,	10		3	Linear	Telefalle	- 55		

Molecular Geometry Activity

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