

Name: _____ **KEY** _____ Period: _____ Date: _____

Determining Molecular Polarity

Molecules are either POLAR or NON-POLAR. This is determined by:

- If all of the bonds are ionic → the compound is just "ionic"
- If all of the bonds are non-polar covalent → then the molecule is non-polar
- If one of the bonds is polar → then the molecule is polar
- If there is more than one polar bond, then look at symmetry
 - If the shape linear, tetrahedron, or trigonal planar and all the bonds are the same, then it had symmetry and is non-polar.
 - If the shape is a pyramid, linear, or bent, then the molecules is NOT symmetrical and is a polar molecule

NPC = Non-Polar Covalent
 PC = Polar Covalent
 I = Ionic

Fill in the following table:

Compound	Lewis Dot Diagram	Bond Type (NPC/PC/I)	Shape	Symmetry (Y/N)	Molecular Polarity (P/NP)
H ₂	$H - H$	H—H (NPC)	Linear	YES	NP
N ₂	$:N \equiv N:$	N—N (NPC)	Linear	YES	NP
O ₂	$:\ddot{O} = \ddot{O}:$	O—O (NPC)	Linear	YES	NP
CaCl ₂	$:\ddot{Cl}: Ca :\ddot{Cl}:$	(I)	No Shape (Ionic)	NONE	IONIC
CO	$:\ddot{C} \equiv \ddot{O}:$	C—O (PC)	Linear	NO	P

Molecular Shape And Polarity Study Guide Answers

**T. W. Graham Solomons, Craig B.
Fryhle, Scott A. Snyder**



Molecular Shape And Polarity Study Guide Answers:

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