

Assessment and Evaluation

ThoughtLab/ ExpressLab/ Investigation	Curriculum Expectations	Assessment Tools/Techniques	Achievement Chart Category	Learning Skills
ExpressLab: A Penny for Your Isotopes, p. 168	Understanding Basic Concepts <ul style="list-style-type: none"> [OCR 1.02] explain the relationship between isotopic abundance and relative atomic mass 	<ul style="list-style-type: none"> Assessment Checklist 5: Learning Skills (see "Assessment and Evaluation" in the front matter of <i>Teacher's Resource CD-ROM</i>) 	<ul style="list-style-type: none"> Knowledge/ Understanding 	<ul style="list-style-type: none"> Teamwork

Solutions to Practice Problems

Student Textbook page 170

See *Solutions Manual* for solutions to Practice Problems.

Section Review Answers

Student Textbook page 170

- The average atomic mass is a weighted average of the atomic masses of the isotopes of potassium. Any given atom would have the atomic mass of one of the isotopes, not the average mass of all the isotopes.

2.

Isotope	Atomic Mass (u)	Percentage Abundance
Mg-24	23.985	78.78
Mg-25	24.985	10.13
Mg-26	25.983	11.17

$$\text{average atomic mass} = \frac{[78.78(23.985) + 10.13(24.985) + 11.17(25.983)] \text{ u}}{100}$$

$$\text{average atomic mass} = \frac{[1887.6 + 253.1 + 290.2] \text{ u}}{100}$$

$$\text{average atomic mass} = 24.31 \text{ u}$$

Therefore, the average atomic mass of magnesium is 24.31 u.

- It is not possible to solve the problem. There would be two unknowns in just one equation. Percentage abundance for one isotope is needed.
- Radioisotopes occur in nature along with the stable isotopes of an element. The nucleus of a radioisotope is continually altered due to the loss of alpha or beta particles. This in turn affects the atomic mass of that isotope as it undergoes radioactive decay. Changing masses such as this should not be taken into consideration when calculating the average atomic mass of an element. Another reason these isotopes are not used in calculating the average atomic mass of silver is that their natural abundance is usually so small that their contribution to the average atomic mass is insignificant.

Mhr Chemistry Review Solutions Unit 5

SB Merriam

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Table of Contents Mhr Chemistry Review Solutions Unit 5

1. Understanding the eBook Mhr Chemistry Review Solutions Unit 5
 - The Rise of Digital Reading Mhr Chemistry Review Solutions Unit 5
 - Advantages of eBooks Over Traditional Books
2. Identifying Mhr Chemistry Review Solutions Unit 5
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Mhr Chemistry Review Solutions Unit 5
 - User-Friendly Interface
4. Exploring eBook Recommendations from Mhr Chemistry Review Solutions Unit 5
 - Personalized Recommendations
 - Mhr Chemistry Review Solutions Unit 5 User Reviews and Ratings
 - Mhr Chemistry Review Solutions Unit 5 and Bestseller Lists

5. Accessing Mhr Chemistry Review Solutions Unit 5 Free and Paid eBooks
 - Mhr Chemistry Review Solutions Unit 5 Public Domain eBooks
 - Mhr Chemistry Review Solutions Unit 5 eBook Subscription Services
 - Mhr Chemistry Review Solutions Unit 5 Budget-Friendly Options
6. Navigating Mhr Chemistry Review Solutions Unit 5 eBook Formats
 - ePub, PDF, MOBI, and More
 - Mhr Chemistry Review Solutions Unit 5 Compatibility with Devices
 - Mhr Chemistry Review Solutions Unit 5 Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Mhr Chemistry Review Solutions Unit 5
 - Highlighting and Note-Taking Mhr Chemistry Review Solutions Unit 5
 - Interactive Elements Mhr Chemistry Review Solutions Unit 5
8. Staying Engaged with Mhr Chemistry Review Solutions Unit 5
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Mhr Chemistry Review Solutions Unit 5
9. Balancing eBooks and Physical Books Mhr Chemistry Review Solutions Unit 5
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Mhr Chemistry Review Solutions Unit 5
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Mhr Chemistry Review Solutions Unit 5
 - Setting Reading Goals Mhr Chemistry Review Solutions Unit 5
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Mhr Chemistry Review Solutions Unit 5
 - Fact-Checking eBook Content of Mhr Chemistry Review Solutions Unit 5
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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