

# CHAPTER 1

## MATTER—ITS PROPERTIES AND MEASUREMENT

### PRACTICE EXAMPLES

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**1A** (E) Convert the Fahrenheit temperature to Celsius and compare.  
 $^{\circ}\text{C} = (^{\circ}\text{F} - 32^{\circ}\text{F}) \frac{5^{\circ}\text{C}}{9^{\circ}\text{F}} = (350^{\circ}\text{F} - 32^{\circ}\text{F}) \frac{5^{\circ}\text{C}}{9^{\circ}\text{F}} = 177^{\circ}\text{C}$ .

**1B** (E) We convert the Fahrenheit temperature to Celsius.  $^{\circ}\text{C} = (^{\circ}\text{F} - 32^{\circ}\text{F}) \frac{5^{\circ}\text{C}}{9^{\circ}\text{F}} = (-15^{\circ}\text{F} - 32^{\circ}\text{F}) \frac{5^{\circ}\text{C}}{9^{\circ}\text{F}} = -26^{\circ}\text{C}$ . The antifreeze only protects to  $-22^{\circ}\text{C}$  and thus it will not offer protection to temperatures as low as  $-15^{\circ}\text{F} = -26.1^{\circ}\text{C}$ .

**2A** (E) The mass is the difference between the mass of the full and empty flask.

$$\text{density} = \frac{291.4 \text{ g} - 108.6 \text{ g}}{125 \text{ mL}} = 1.46 \text{ g/mL}$$

**2B** (E) First determine the volume required.  $V = (1.000 \times 10^3 \text{ g}) \div (8.96 \text{ g cm}^{-3}) = 111.6 \text{ cm}^3$ . Next determine the radius using the relationship volume of a sphere and radius.

$$V = \frac{4}{3} \pi r^3 = 111.6 \text{ cm}^3 = \frac{4}{3} (3.1416) r^3 \quad r = \sqrt[3]{\frac{111.6 \times 3}{4(3.1416)}} = 2.987 \text{ cm}$$

**3A** (E) The volume of the stone is the difference between the level in the graduated cylinder with the stone present and with it absent.

$$\text{density} = \frac{\text{mass}}{\text{volume}} = \frac{28.4 \text{ g rock}}{44.1 \text{ mL rock \& water} - 33.8 \text{ mL water}} = 2.76 \text{ g/mL} = 2.76 \text{ g/cm}^3$$

**3B** (E) The water level will remain unchanged. The mass of the ice cube displaces the same mass of liquid water. A 10.0 g ice cube will displace 10.0 g of water. When the ice cube melts, it simply replaces the displaced water, leaving the liquid level unchanged.

**4A** (E) The mass of ethanol can be found using dimensional analysis.

$$\begin{aligned} \text{ethanol mass} &= 25 \text{ L gasohol} \times \frac{1000 \text{ mL}}{1 \text{ L}} \times \frac{0.71 \text{ g gasohol}}{1 \text{ mL gasohol}} \times \frac{10 \text{ g ethanol}}{100 \text{ g gasohol}} \times \frac{1 \text{ kg ethanol}}{1000 \text{ g ethanol}} \\ &= 1.8 \text{ kg ethanol} \end{aligned}$$

**4B** (E) We use the mass percent to determine the mass of the 25.0 mL sample.

$$\text{rubbing alcohol mass} = 15.0 \text{ g (2-propanol)} \times \frac{100.0 \text{ g rubbing alcohol}}{70.0 \text{ g (2-propanol)}} = 21.43 \text{ g rubbing alcohol}$$

$$\text{rubbing alcohol density} = \frac{21.4 \text{ g}}{25.0 \text{ mL}} = 0.857 \text{ g/mL}$$

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**Mohammad Mansoob Khan**



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without having to rely much on its mathematical aspect making it more approachable to those who are new to quantum chemistry The book covers topics such as orbitals in quantum chemical calculations electronic ionizations and transitions molecular orbital charge distributions orbital transformations and calculations not involving orbitals and electron densities and shapes in atoms and molecules Also included in the book are the cross sectional plots of electron densities of compounds such as organic compounds like methane ethane and ethylene monomeric lithium fluoride and monomeric methyl lithium hydrogen cyanide and methinophosphide and monomeric borane and diborane The text is recommended for those who have begun taking an interest in quantum chemistry but do not wish to deal yet with the mathematics part of the subject

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