

CHAPTER 1

MATTER—ITS PROPERTIES AND MEASUREMENT

PRACTICE EXAMPLES

- 1A** (E) Convert the Fahrenheit temperature to Celsius and compare.
 $^{\circ}\text{C} = (^{\circ}\text{F} - 32^{\circ}\text{F}) \frac{5^{\circ}\text{C}}{9^{\circ}\text{F}} = (350^{\circ}\text{F} - 32^{\circ}\text{F}) \frac{5^{\circ}\text{C}}{9^{\circ}\text{F}} = 177^{\circ}\text{C}.$

- 1B** (E) We convert the Fahrenheit temperature to Celsius.
 $^{\circ}\text{C} = (^{\circ}\text{F} - 32^{\circ}\text{F}) \frac{5^{\circ}\text{C}}{9^{\circ}\text{F}} = (-15^{\circ}\text{F} - 32^{\circ}\text{F}) \frac{5^{\circ}\text{C}}{9^{\circ}\text{F}} = -26^{\circ}\text{C}.$ The antifreeze only protects to -22°C and thus it will not offer protection to temperatures as low as $-15^{\circ}\text{F} = -26.1^{\circ}\text{C}.$

- 2A** (E) The mass is the difference between the mass of the full and empty flask.

$$\text{density} = \frac{291.4 \text{ g} - 108.6 \text{ g}}{125 \text{ mL}} = 1.46 \text{ g/mL}$$

- 2B** (E) First determine the volume required. $V = (1.000 \times 10^3 \text{ g}) \div (8.96 \text{ g cm}^{-3}) = 111.6 \text{ cm}^3.$ Next determine the radius using the relationship between volume of a sphere and radius.

$$V = \frac{4}{3} \pi r^3 = 111.6 \text{ cm}^3 = \frac{4}{3} (3.1416) r^3 \quad r = \sqrt[3]{\frac{111.6 \times 3}{4(3.1416)}} = 2.987 \text{ cm}$$

- 3A** (E) The volume of the stone is the difference between the level in the graduated cylinder with the stone present and with it absent.

$$\text{density} = \frac{\text{mass}}{\text{volume}} = \frac{28.4 \text{ g rock}}{44.1 \text{ mL rock \& water} - 33.8 \text{ mL water}} = 2.76 \text{ g/mL} = 2.76 \text{ g/cm}^3$$

- 3B** (E) The water level will remain unchanged. The mass of the ice cube displaces the same mass of liquid water. A 10.0 g ice cube will displace 10.0 g of water. When the ice cube melts, it simply replaces the displaced water, leaving the liquid level unchanged.

- 4A** (E) The mass of ethanol can be found using dimensional analysis.

$$\begin{aligned} \text{ethanol mass} &= 25 \text{ L gasohol} \times \frac{1000 \text{ mL}}{1 \text{ L}} \times \frac{0.71 \text{ g gasohol}}{1 \text{ mL gasohol}} \times \frac{10 \text{ g ethanol}}{100 \text{ g gasohol}} \times \frac{1 \text{ kg ethanol}}{1000 \text{ g ethanol}} \\ &= 1.8 \text{ kg ethanol} \end{aligned}$$

- 4B** (E) We use the mass percent to determine the mass of the 25.0 mL sample.

$$\text{rubbing alcohol mass} = 15.0 \text{ g (2-propanol)} \times \frac{100.0 \text{ g rubbing alcohol}}{70.0 \text{ g (2-propanol)}} = 21.43 \text{ g rubbing alcohol}$$

$$\text{rubbing alcohol density} = \frac{21.4 \text{ g}}{25.0 \text{ mL}} = 0.857 \text{ g/mL}$$

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Mohammad Mansoob Khan



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