# Unit 3 Energy Changes and Rates of Reaction

#### Solutions to Practice Problems in Chapter 5 Energy Changes

# Calculating the Absorption of Heat (Student textbook page 281)

 How much heat must be added to a 36.2 g sample of iron to increase its temperature by 250.0°C?

#### What Is Required?

You have to calculate the amount of heat that must be added to a sample of iron to increase its temperature by 250.0°C.

#### What Is Given?

You know the mass of iron: m = 36.2 g

You know the change in the temperature:  $\Delta T = 250.0^{\circ}$ C

You have the specific heat capacity of iron from Table 5.1 on page 280 of the student text:  $c_{Fe} = 0.449 \text{ J/g} \cdot ^{\circ}\text{C}$ 

Plan Your Strategy	Act on Your Strategy
Substitute the given information into the formula $Q = mc\Delta T$ . Solve for $Q$ .	$Q = mc\Delta T$ = $(36.2 \text{ f}) \left(0.449 \frac{\text{J}}{\text{f} \cdot 26}\right) (250.026)$ = $4063.45 \text{ J}$ = $4.06 \text{ kJ}$ The iron must gain $4060 \text{ J}$ or $4.06 \text{ kJ}$ of heat energy.

#### Check Your Solution

Check that the information has been substituted correctly. The answer has the correct number of significant digits.

# **Mhr Chemistry 12 Solutions**

**Emilie Sanchez** 

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