#### WORKSHEET - ASSIGNING OXIDATION NUMBERS

Name \_\_\_\_\_ Period \_\_\_\_

#### Oxidation Number Rules:

1. A pure element has an oxidation number of 0.

The exidation number of an element in a monatomic ion equals the charge of the ion.

The sum of the oxidation number of all the elements in a compound equals 0.

 The sum of the oxidation numbers of all the elements in a polyatomic ion equals the charge on the ion.

Directions: In the following questions, give the oxidation number of the indicated atoms/ion.

N in N<sub>2</sub>O<sub>3</sub> N<sup>3\*</sup>

2. S in H<sub>2</sub>SO<sub>4</sub> S

3. C C\*

4. C in CO C\*\*

Na in NaCl Na\*

6 Hin H<sub>2</sub>O H

7. Ba in BaCl<sub>2</sub> Ba<sup>3\*</sup>

8 N in NO<sub>2</sub> N<sup>9\*</sup>

9. S in Al<sub>2</sub>S<sub>0</sub> S<sup>3</sup>

10. S in HSO. St

Cl in Fe(ClO<sub>2</sub>)<sub>3</sub> Cf<sup>\*\*</sup>

Fe in Fe(ClO<sub>2</sub>)<sub>3</sub> Fe<sup>3\*</sup>

13. N in NO<sub>3</sub> No

14. Cu2\* Cu2\*

15. Zn2+ Zn2+

16. C in CH, C"

17. Min in MinO2. Min#

18. S in SO<sub>2</sub>2- S\*\*

19. Mg2" Mg3"

20. Cl CI

21. Oz Oz\*

22 P. P.

23. Na in Na<sub>2</sub>S Na

24 S in H<sub>2</sub>S 5°

25. Ca\*\* Ca\*\*

26 C in CN C\*\*

27. H in OH H

28. Mn in KMnO. Mn\*\*

29. I in Mg(IO<sub>2</sub>)<sub>2</sub> I\*\*

30. C in C<sub>2</sub>O<sub>4</sub><sup>2</sup> C<sup>2\*</sup>

# **Oxidation Numbering Problems And Answers**

**Camilla Rothe** 

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