#### Reduction-Oxidation Reactions Worksheet

## Part I: Identifying parts of Simple Red-ox reactions

For the equations below, all in aqueous solution:

- 1) assign oxidation numbers to each atom
- identify the reactant oxidized,
- 3) the reactant reduced,
- the oxidizing agent,
- 5) the reducing agent,
- write the balanced oxidation and reduction half reactions.

Example: oxidized reduced

Mg° + Br₂° → Mg\*² Br¹²

Reducing agent oxidizing agent

Oxidation half reaction: Mg° → Mg\*² + 2 e°

Reduction half reaction: 2 e° + Br₂° → 2Br°

Fe + Zn<sup>+2</sup> → Fe<sup>+2</sup> + Zn

Al + Fe<sup>+2</sup> → Al<sup>+3</sup> + Fe

3. Cu + AgNO<sub>3</sub>  $\rightarrow$  Cu(NO<sub>3</sub>)<sub>2</sub> + Ag

SiCl<sub>4</sub> (I) + Mg (s) → MgCl<sub>2</sub> (s) + Si (s)

## **Oxidation Reduction Reactions Study Guide Answers**

**Arshad Iqbal** 

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