

# HIGHER SCHOOL CERTIFICATE EXAMINATION Chemistry

## DATA SHEET

Avogadro constant, $N_A$ .....	$6.022 \times 10^{23} \text{ mol}^{-1}$
Volume of 1 mole ideal gas: at 100 kPa and	
at 0°C (273.15 K) .....	22.71 L
at 25°C (298.15 K) .....	24.79 L
Ionisation constant for water at 25°C (298.15 K), $K_w$ .....	$1.0 \times 10^{-14}$
Specific heat capacity of water .....	$4.18 \times 10^3 \text{ J kg}^{-1} \text{ K}^{-1}$

### Some useful formulae

$$\text{pH} = -\log_{10}[\text{H}^+]$$

$$\Delta H = -m C \Delta T$$

### Some standard potentials

$\text{K}^+ + \text{e}^-$	$\text{e}^0$	$\text{K}(\text{s})$	-2.94 V
$\text{Ba}^{2+} + 2\text{e}^-$	$\text{e}^0$	$\text{Ba}(\text{s})$	-2.91 V
$\text{Ca}^{2+} + 2\text{e}^-$	$\text{e}^0$	$\text{Ca}(\text{s})$	-2.87 V
$\text{Na}^+ + \text{e}^-$	$\text{e}^0$	$\text{Na}(\text{s})$	-2.71 V
$\text{Mg}^{2+} + 2\text{e}^-$	$\text{e}^0$	$\text{Mg}(\text{s})$	-2.36 V
$\text{Al}^{3+} + 3\text{e}^-$	$\text{e}^0$	$\text{Al}(\text{s})$	-1.68 V
$\text{Mn}^{2+} + 2\text{e}^-$	$\text{e}^0$	$\text{Mn}(\text{s})$	-1.18 V
$\text{H}_2\text{O} + \text{e}^-$	$\text{e}^0$	$\frac{1}{2}\text{H}_2(\text{g}) + \text{OH}^-$	-0.83 V
$\text{Zn}^{2+} + 2\text{e}^-$	$\text{e}^0$	$\text{Zn}(\text{s})$	-0.76 V
$\text{Fe}^{2+} + 2\text{e}^-$	$\text{e}^0$	$\text{Fe}(\text{s})$	-0.44 V
$\text{Ni}^{2+} + 2\text{e}^-$	$\text{e}^0$	$\text{Ni}(\text{s})$	-0.24 V
$\text{Sn}^{2+} + 2\text{e}^-$	$\text{e}^0$	$\text{Sn}(\text{s})$	-0.14 V
$\text{Pb}^{2+} + 2\text{e}^-$	$\text{e}^0$	$\text{Pb}(\text{s})$	-0.13 V
$\text{H}^+ + \text{e}^-$	$\text{e}^0$	$\frac{1}{2}\text{H}_2(\text{g})$	0.00 V
$\text{SO}_4^{2-} + 4\text{H}^+ + 2\text{e}^-$	$\text{e}^0$	$\text{SO}_2(\text{aq}) + 2\text{H}_2\text{O}$	0.16 V
$\text{Cu}^{2+} + 2\text{e}^-$	$\text{e}^0$	$\text{Cu}(\text{s})$	0.34 V
$\frac{1}{2}\text{O}_2(\text{g}) + \text{H}_2\text{O} + 2\text{e}^-$	$\text{e}^0$	$2\text{OH}^-$	0.40 V
$\text{Cu}^+ + \text{e}^-$	$\text{e}^0$	$\text{Cu}(\text{s})$	0.52 V
$\frac{1}{2}\text{I}_2(\text{s}) + \text{e}^-$	$\text{e}^0$	$\text{I}^-$	0.54 V
$\frac{1}{2}\text{I}_2(\text{aq}) + \text{e}^-$	$\text{e}^0$	$\text{I}^-$	0.62 V
$\text{Fe}^{3+} + \text{e}^-$	$\text{e}^0$	$\text{Fe}^{2+}$	0.77 V
$\text{Ag}^+ + \text{e}^-$	$\text{e}^0$	$\text{Ag}(\text{s})$	0.80 V
$\frac{1}{2}\text{Br}_2(\text{l}) + \text{e}^-$	$\text{e}^0$	$\text{Br}^-$	1.08 V
$\frac{1}{2}\text{Br}_2(\text{aq}) + \text{e}^-$	$\text{e}^0$	$\text{Br}^-$	1.10 V
$\frac{1}{2}\text{O}_2(\text{g}) + 2\text{H}^+ + 2\text{e}^-$	$\text{e}^0$	$\text{H}_2\text{O}$	1.23 V
$\frac{1}{2}\text{Cl}_2(\text{g}) + \text{e}^-$	$\text{e}^0$	$\text{Cl}^-$	1.36 V
$\frac{1}{2}\text{Cr}_2\text{O}_7^{2-} + 7\text{H}^+ + 3\text{e}^-$	$\text{e}^0$	$\text{Cr}^{3+} + \frac{7}{2}\text{H}_2\text{O}$	1.36 V
$\frac{1}{2}\text{Cl}_2(\text{aq}) + \text{e}^-$	$\text{e}^0$	$\text{Cl}^-$	1.40 V
$\text{MnO}_4^- + 8\text{H}^+ + 5\text{e}^-$	$\text{e}^0$	$\text{Mn}^{2+} + 4\text{H}_2\text{O}$	1.51 V
$\frac{1}{2}\text{F}_2(\text{g}) + \text{e}^-$	$\text{e}^0$	$\text{F}^-$	2.89 V

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**LP Steffe**



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