

Chemistry – Unit 11 – Worksheet 2

Periodic trends

1. On the graph of the first ionization energy (IE₁) vs. the atomic number (Z) for the first 20 elements, label each point with the symbol of the element it represents.
2. On the graph of the atomic radius (in nm) vs. the atomic number, label each point with the symbol of the element it represents.

Group = _____

Period = _____

Conclusion

1. Examine the graph of 1st ionization energy vs. atomic number. Can a periodic trend be observed? If so, describe the trend.
2. Which group (or column #) has the highest ionization energies? Which group has the lowest?
3. Examine the graph of atomic radius vs. atomic number. Can a periodic trend be observed? If so, describe the trend.
4. For a given period (a row in the Periodic Table), which group (a column in the table) appears to have the largest atomic radii? Which group has the smallest?
5. What appears to be the relationship between trends in IE and atomic radii?

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