

Topic	Grade	What do you need to know	 or 	What do You have to work on?
Atomic Structure	A	<ul style="list-style-type: none"> <li>Why an atom is neutral in terms of sub-atomic particles</li> <li>Outline isotopes from data about numbers of protons and neutrons</li> </ul>		
	B	<ul style="list-style-type: none"> <li>Know the charges and masses of protons, electrons and neutrons.</li> <li>Know how isotopes are the same but have different mass number.</li> </ul>		
	C	<ul style="list-style-type: none"> <li>Describe the arrangement of an atoms electron structure - 2.8.8.2</li> <li>Describe the arrangement of elements in the periodic table.</li> </ul>		
Ionic Bonding	A	<ul style="list-style-type: none"> <li>Explain ionic bonding using the 'dot and cross' model.</li> <li>Explain the melting point of NaCl by the structure of and bonding.</li> <li>Explain why the melting point of NaCl is lower than that of MgO.</li> </ul>		
	B	<ul style="list-style-type: none"> <li>Predict the properties of substances with giant ionic structures.</li> <li>Understand that atoms with a full outer shell are electronically stable.</li> </ul>		
	C	<ul style="list-style-type: none"> <li>Ionic bonds are formed from metals and non-metals</li> <li>Electrons are donated and received.</li> <li>Metals always donate the electron(s).</li> <li>Non- metals always receive the electron(s)</li> </ul>		
The Periodic Table	A	<ul style="list-style-type: none"> <li>Explain some covalent bonds using a 'dot and cross' model</li> <li>Explain double covalent binds using a 'dot and cross' model</li> </ul>		
	B	<ul style="list-style-type: none"> <li>Know why CO<sub>2</sub> and H<sub>2</sub>O have low melting points</li> </ul>		
	C	<ul style="list-style-type: none"> <li>Covalent bonds always have non-metals and share electrons</li> <li>Know that the group number is the same as the number of electrons in the outer shell.</li> <li>Be able to find out how many shells an atom has from the period it is in.</li> </ul>		
The Group 1 Elements	A	<ul style="list-style-type: none"> <li>Construct a balanced symbol equation for Cu<sup>2+</sup>, Fe<sup>2+</sup> and Fe<sup>3+</sup> to make Cu(OH)<sub>2</sub>, Fe(OH)<sub>2</sub>, Fe(OH)<sub>3</sub>.</li> <li>Reactivity with water gives NaOH and KOH.</li> <li>Know the process is oxidation.</li> </ul>		
	B	<ul style="list-style-type: none"> <li>Why group 1 elements form similar positive ions.</li> <li>Construct a balance symbol equation of group 1 with water</li> </ul>		
	C	<ul style="list-style-type: none"> <li>Metal + Water → Metal Hydroxide + Hydrogen</li> <li>Know how to use a flame test to identify potassium, lithium and sodium compounds.</li> </ul>		
The Group 7 Elements	A	<ul style="list-style-type: none"> <li>Predict if displacement reactions will occur with other halogens</li> <li>Explain the trend of reactivity of group 7 by ideas of electron gain</li> <li>Explain why a process is reduced from its ionic</li> </ul>		

# Ocr Gateway Separate Sciences B C4 Checklist

**Yi-Tong Ma**



## **Ocr Gateway Separate Sciences B C4 Checklist:**

## **Ocr Gateway Separate Sciences B C4 Checklist** Book Review: Unveiling the Power of Words

In a world driven by information and connectivity, the ability of words has become more evident than ever. They have the capability to inspire, provoke, and ignite change. Such could be the essence of the book **Ocr Gateway Separate Sciences B C4 Checklist**, a literary masterpiece that delves deep into the significance of words and their affect our lives. Written by a renowned author, this captivating work takes readers on a transformative journey, unraveling the secrets and potential behind every word. In this review, we shall explore the book's key themes, examine its writing style, and analyze its overall effect on readers.

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