



Spectroscopy Lab

Background Info:

Bohr proposed that an electron is found only in specific circular paths, or orbits, around the nucleus. As a result, each orbit has a fixed energy called an energy level. These energy levels (orbits) are like the rungs of a ladder (See Figure A). The electrons cannot be found in between these energy levels, just like a person cannot stand in between the rungs. Electrons must gain energy (become excited) to move up energy levels. Electrons can become excited by heat, light, electricity etc. These high energy electrons are unstable and eventually fall back down to their lower energy levels (ground state) releasing the energy that they had gained when they were initially excited. This energy is released in the form of light and is in what Bohr observed.

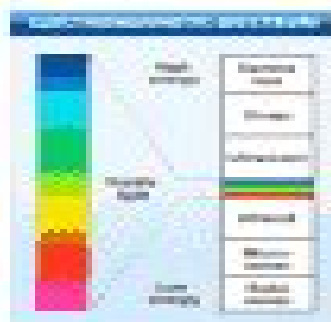
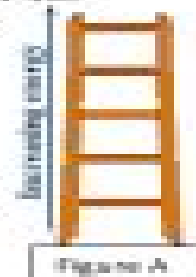


Figure B

White light will produce a continuous spectrum when viewed using a prism or spectroscope (see Figure B). Different colors of light have different energies. When electrons of an atom are excited they will release different colors of light as they fall back to ground state (this is called an emission line spectra). The colors of light correspond to the amount of energy released (See Figure C). Using a spectroscope you can view the emission line spectra. Different elements produce different spectra that are unique enough to be considered a "fingerprint" of the element.



Figure C

Objective:

In this virtual lab you will:

1. Observe the bright line spectra (emission spectra) for various elements.
2. Use a flame test to observe the color produced when metal ions are heated.
3. Identify unknown metal ions based on the results of the flame test.

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