

14**THE BEHAVIOR OF GASES****Chapter Test A****A. Matching**

Match each description in Column B with the correct term in Column A. Write the letter of the correct description on the line.

Column A**Column B**

- | | |
|--|--|
| _____ 1. ideal gas constant (R) | a. The volume of a fixed mass of gas is directly proportional to the Kelvin temperature if the volume is kept constant. |
| _____ 2. Boyle's law | b. At constant volume and temperature, the total pressure exerted by a mixture of gases is equal to the sum of the partial pressures of the component gases. |
| _____ 3. Dalton's law of partial pressures | c. $8.31 \frac{\text{L} \cdot \text{kPa}}{\text{K} \cdot \text{mol}}$ |
| _____ 4. ideal gas law | d. the contribution each gas in a mixture makes to the total pressure |
| _____ 5. combined gas law | e. A gas tends to move to an area of lower concentration until the concentration is uniform throughout. |
| _____ 6. Charles's law | f. $\frac{P_1 \times V_1}{T_1} = \frac{P_2 \times V_2}{T_2}$ |
| _____ 7. diffusion | g. $P \times V = n \times R \times T$ |
| _____ 8. partial pressure | h. For a given mass of gas at constant temperature, the volume of gas varies inversely with the pressure. |

B. Multiple Choice

Choose the best answer and write its letter on the line.

- | | | |
|--|--------------------|------------------|
| _____ 9. As the temperature of a fixed volume of gas increases, the pressure will | a. vary inversely. | c. be unchanged. |
| | b. decrease. | d. increase. |
| _____ 10. A breathing mixture used by deep-sea divers contains helium, oxygen, and carbon dioxide. What is the partial pressure of oxygen at 101.3 kPa total pressure if $P_{\text{He}} = 84.0 \text{ kPa}$ and $P_{\text{CO}_2} = 0.10 \text{ kPa}$? | a. 10.3 kPa | c. 34.4 kPa |
| | b. 17.2 kPa | d. 185.4 kPa |

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design courses case studies and design projects with practical applications and appendixes with current equipment cost data and preliminary design information for eleven chemical processes including seven brand new to this edition

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