

4.66. Methanol is formed from carbon monoxide and hydrogen in the gas-phase reaction



(A) (B) (C)

The mole fractions of the reactive species at equilibrium satisfy the relation

$$\frac{y_C}{y_A y_B^2} P^2 = K_e(T)$$

where P is the total pressure (atm), K_e the reaction equilibrium constant (atm^{-2}), and T the temperature (K). The equilibrium constant K_e equals 10.5 at 373 K , and 2.316×10^{-4} at 573 K . A semilog plot of K_e (logarithmic scale) versus $1/T$ (rectangular scale) is approximately linear between $T = 300\text{ K}$ and $T = 600\text{ K}$.

- Derive a formula for $K_e(T)$, and use it to show that $K_e(450\text{ K}) = 0.0548\text{ atm}^{-2}$.
- Write expressions for n_A , n_B , and n_C (gram-moles of each species), and then y_A , y_B , and y_C , in terms of n_{A0} , n_{B0} , n_{C0} , and ξ , the extent of reaction. Then derive an equation involving only n_{A0} , n_{B0} , n_{C0} , P , T , and ξ , where ξ is the extent of reaction at equilibrium.
- Suppose you begin with equimolar quantities of CO and H_2 and no CH_3OH , and the reaction proceeds to equilibrium at 423 K and 2.00 atm . Calculate the molar composition of the product (y_A , y_B , and y_C) and the fractional conversion of CO .

Question And Answers Chemical Processes And Principles

**Fred Herbert Colvin, Lester Gray
French**



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STOICHIOMETRY AND PROCESS CALCULATIONS K.V. NARAYANAN ,B. LAKSHMIKUTTY,2016-12-01 Designed as a textbook for the undergraduate students of chemical engineering and related disciplines such as biotechnology polymer technology petrochemical engineering electrochemical engineering environmental engineering and safety engineering the chief objective of the book is to prepare students to make analysis of chemical processes through calculations and to develop systematic problem solving skills in them The text presents the fundamentals of chemical engineering operations and processes in a simple style that helps the students to gain a thorough understanding of chemical process calculations The book deals with the principles of stoichiometry to formulate and solve material and energy balance problems in processes with and without chemical reactions With the help of examples the book explains the construction and use of reference substance plots equilibrium diagrams psychrometric charts steam tables and enthalpy composition diagrams It also elaborates on thermophysics and thermochemistry to acquaint the students with the thermodynamic principles of energy balance calculations The book is supplemented with Solutions Manual for instructors containing detailed solutions of all chapter end unsolved problems NEW TO THE SECOND EDITION Incorporates a new chapter on Bypass Recycle and Purge Operations Comprises updations in some sections and presents new sections on Future Avenues and Opportunities in Chemical Engineering Processes in Biological and Energy Systems Contains several new worked out examples in the chapter on Material Balance with Chemical Reaction Includes GATE questions with answers up to the year 2016 in Objective type questions KEY FEATURES SI units are used throughout the book All basic chemical engineering operations and processes are introduced and different types of problems are illustrated with worked out examples Stoichiometric principles are extended to solve problems related to bioprocessing environmental engineering etc Exercise problems more than 810 are organised according to the difficulty level and all are provided with answers **Green Chemistry and Sustainable**

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