

1. In the list below, circle the properties of acids and box the properties of bases.

Sour taste	Feel slippery	Contains hydroxide ions
Turns litmus blue	Reacts with metals to form H_2	Turns litmus red
Bitter taste	Contains hydrogen ions	Corrosive

2. Complete the chart below:

Acid-Base Definitions		
Type	Acid Definition	Base Definition
Arrhenius	Donates an H^+ ion in solution	Donates an OH^- ion in solution
Bronsted-Lowry	Donates an H^+ ion	Accepts an H^+ ion
Lewis	Accepts an electron pair	Donates an electron pair

3. What term(s) describe a substance that can act as an acid or a base? **Amphoteric or amphiprotic**
4. Neutralization reactions always produce a **SALT** and **WATER**
5. If the $[H^+]$ is greater than the $[OH^-]$, the solution is acidic basic neutral
6. If the $[OH^-]$ is greater than the $[H^+]$, the solution is acidic basic neutral
7. If the $[H^+]$ is equal to the $[OH^-]$, the solution is acidic basic neutral
8. If the pH is 7, the solution is acidic basic neutral
9. If the pH is greater than 7, the solution is acidic basic neutral
10. If the pH is less than 7, the solution is acidic basic neutral
11. How many moles of lithium hydroxide are required to neutralize 19.2 grams of chloric acid?



$$(19.2 \text{ g } HClO_3) \times (1 \text{ mol } HClO_3 / 84 \text{ g } HClO_3) \times (1 \text{ mol } LiOH / 1 \text{ mol } HClO_3) = \underline{0.228 \text{ mol } LiOH}$$

12. How many grams of calcium hydroxide are required to neutralize 132.4 grams of nitric acid?



$$(132.4 \text{ g } HNO_3) \times (1 \text{ mol } HNO_3 / 63 \text{ g } HNO_3) \times (1 \text{ mol } Ca(OH)_2 / 2 \text{ mol } HNO_3) \times (74 \text{ g } Ca(OH)_2 / 1 \text{ mol } Ca(OH)_2) = \underline{77.76 \text{ g } Ca(OH)_2}$$

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L Cohen



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Basic Analytical Chemistry L. Pataki,E. Zapp,2013-10-22 Pergamon Series in Analytical Chemistry Volume 2 Basic Analytical Chemistry brings together numerous studies of the vast expansion in the use of classical and instrumental methods of analysis This book is composed of six chapters After providing a theoretical background of analytical chemistry this book goes on dealing with the fundamental principles of chemical equilibria in solution The subsequent chapters consider the advances in qualitative and quantitative chemical analyses These chapters present a unified view of these analyses based on the Bronsted Lowry theory and the donor acceptor principle These topics are followed by discussions on instrumental analysis using various methods including electrochemical optical spectroscopic and thermal methods as well as radioactive isotopes The final chapters examine the separation methods and the essential features of organic chemical analysis that are different from methods for inorganic compounds This book is of value to analytical chemists and researchers **Hard and**

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Soft Matter at Aqueous Interfaces Peter Lang,Yi Liu,2015-12-24 This book covers the science of interfaces between an aqueous phase and a solid another liquid or a gaseous phase starting from the basic physical chemistry all the way to state of the art research developments Both experimental and theoretical methods are treated thanks to the contributions of a distinguished list of authors who are all active researchers in their respective fields The properties of these interfaces are crucial for a wide variety of processes products and biological systems and functions such as the formulation of personal care and food products paints and coatings microfluidic and lab on a chip applications cell membranes and lung surfactants Accordingly research and expertise on the subject are spread over a broad range of academic disciplines and industrial laboratories This book brings together knowledge from these different places with the aim of fostering education collaborations and research progress

Chemical Reactions in Solvents and Melts G. Charlot,B. Trémillon,2013-10-22 Chemical Reactions in Solvents and Melts discusses the use of organic and inorganic compounds as well as of melts as solvents This book examines the applications in organic and inorganic chemistry as well as in electrochemistry Organized into two parts encompassing 15 chapters this book begins with an overview of the general properties and the different types of reactions including acid base reactions complex formation reactions and oxidation reduction reactions This text then describes the properties of inert and active solvents Other chapters consider the proton transfer reactions in polar solvents as well as the transfer of other ions This book discusses as well the solubility in a number of solvents by the formation of different bonds between the solute and the solvent molecule The final chapter deals with the general characteristics of the oxidation reduction reactions of melts This book is a valuable resource for chemists students and researchers The American Dispensatory, John Redman Coxe,1830 *A New & Complete Dictionary of Arts & Sciences*, George Gregory,1815

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