

Colligative properties are properties that depend only upon the number of solute atoms, ions, or molecules in a solution and not on the nature of those atoms, ions or molecules. Freezing point depression and boiling point elevation are examples of colligative properties. Raoult discovered that the addition of solute particles causes the boiling point of a solution to be elevated and the freezing point to be depressed. It has been found that 1 mole of nonvolatile, nonionizing solute particles will raise the boiling point of 1.0 kg of water by 0.52 °C, to 100.52°C at 1 atm of pressure. The same concentration will lower the freezing point of 1.0 kg of water by 1.86°C, to -1.86°C. These two numbers apply only to water and are called the molal boiling point and freezing point constants for water. Other solvents have different constants.

PROPERTY	BOILING POINT ELEVATION	FREEZING POINT DEPRESSION
EQUATION	Boiling Point Elevation: A solution will boil at a higher temperature than the pure solvent. $\Delta T_b = i K_b m$ ΔT_b = change in boiling temp. in (°C) <i>i</i> = a unitless constant (called the van't Hoff factor) associated with the degree of dissociation of the solute in a solvent. K_b = molal boiling point elevation constant (°C/m) For H₂O: K_b = .52°C/m <i>m</i> = molality of the solution	Freezing Point Depression: A solution will freeze at a lower temperature than the pure solvent. $\Delta T_f = i K_f m$ ΔT_f = change in freezing temp. in (°C) <i>i</i> = a unitless constant (called the van't Hoff factor) associated with the degree of dissociation of the solute in a solvent. K_f = molal freezing point lowering constant (°C/m) For H₂O: K_f = 1.86°C/m <i>m</i> = molality of the solution
	<i>i</i> = 1 for substances that do not ionize 2 for substances which ionize into two ions like LiCl 3 for substances which ionize into three ions like CaCl ₂ ...etc. <i>i</i> may have to be calculated if a % of dissociation is given for a solute. If the solvent is not water you must find its K_b .	<i>i</i> = 1 for substances that do not ionize 2 for substances which ionize into two ions like LiCl 3 for substances which ionize into three ions like CaCl ₂ ...etc. <i>i</i> may have to be calculated if a % of dissociation is given for a solute. If the solvent is not water you must find its K_f .

EXAMPLE #1

If 90.0 g of nonionizing C₆H₁₂O₆, are dissolved in 255 g of H₂O, what will be the boiling point of the resulting solution?

First calculate the molality (*m*) of the solution:

$$m = \frac{92g}{\left(180 \frac{g}{mole}\right)(0.255 \text{ kg of solvent})} = 2.0 \text{ m}$$

Next apply the boiling point elevation equation:

[*i* = 1 for a nonionizing solute.]

$$\Delta T_b = i K_b m$$

$$\Delta T_b = (1)(0.52^\circ\text{C/m})(2.0 \text{ m})$$

$$\Delta T_b = 1.04^\circ\text{C}$$

$$BP = (\Delta T_b + 100.00^\circ\text{C})$$

$$BP = (1.04^\circ\text{C} + 100.00^\circ\text{C}) = \underline{101.04^\circ\text{C}}$$

Modern Chemistry Study Guide Colligative Properties Answers

Hugh Stott Taylor



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Polymeric Foams Structure-Property-Performance Bernard Obi, 2017-12-07 Polymeric Foams Structure Property Performance A Design Guide is a response to the design challenges faced by engineers in a growing market with evolving standards new regulations and an ever increasing variety of application types for polymeric foam Bernard Obi an author with wide experience in testing characterizing and applying polymer foams approaches this emerging complexity with a practical design methodology that focuses on understanding the relationship between structure properties of polymeric foams and their performance attributes The book not only introduces the fundamentals of polymer and foam science and engineering but also goes more in depth covering foam processing properties and uses for a variety of applications By connecting the diverse technologies of polymer science to those from foam science and by linking both micro and macrostructure property relationships to key performance attributes the book gives engineers the information required to solve pressing design problems involving the use of polymeric foams and to optimize foam performance With a focus on applications in the automotive and transportation industries as well as uses of foams in structural composites for lightweight applications the author provides numerous case studies and design examples of real life industrial problems from various industries and their solutions Provides the science and engineering fundamentals relevant for solving polymer foam application problems Offers

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