

is not necessarily important to give the order of reaction.

- vii. **The order of reaction is obtained from the rate expression but the rate expression is obtained from experiment justify.**

There is no simple relationship b/w the overall reaction and the rate law expression. A reaction may consist of several elementary steps(reactions) and the rate law expression is the combined result of these steps. That is why, we cannot predict the rate law expression by looking at the overall equation. Therefore, the rate law expression must be determined experimentally.

- viii. **Differentiate between fast step and rate determining steps.**

Sr #	Fast step	Rate determining step
1	An elementary reaction that is fast in the reaction mechanism is called fast step.	The slowest step in the reaction mechanism is called rate determining step.
2	It does not control the rate of reaction.	It controls the rate of reaction.

- ix. **50% of hypothetical 1st order reaction completes in one hour ,the remaining 50 percent needs more than that one hour to complete.**

The half life for the first order reaction is the time required to convert 50% of the reactants to products. This occurs at 1 hour, so, $t_{1/2} = 1$ hour. It means in one half-life (1 hour) the concentration decreases 50% and 50% of reactants will be left behind. At the end of second half life, which occurs at 2 hours, the concentration decreases by one half again from 50% to $50\% \div 2 = 25\%$. It means 75% of the reactants are completed in two hours and 25% would be left behind and so on. Thus the remaining 50% needs more than 1 hour to complete reaction.

- x. **The radioactive decay is always a 1st order reaction. Why?**

The half-life of a radioactive substance is the time required for disintegration of half of the element into daughter elements. The half life of a radioactive substance is independent of the amount of the sample. For the first order reaction, the half life is also independent of the initial concentration of the reactant. Therefore, the radioactive decay is always the 1st order of reaction.

- xi. **Differentiate between enthalpy change of reaction and energy of activation of the reaction?**

Sr #	Enthalpy change of a reaction	Energy of activation of reaction
1.	Enthalpy change equals the heat of reaction at constant pressure.	The minimum amount of energy in addition to average K.E which the particles must have for the effective collisions is called activation energy.
2.	$\Delta H = q_p$	$E_a = -2.303 RT = \text{slope}$
3.	Enthalpy change depends on the initial and final state of the reaction.	Activation energy depends upon the slope.

- xii. **Differentiate between homogeneous and heterogeneous catalysts?**

Sr #	Homogeneous catalysts	Heterogeneous catalysts
1	In this catalysis, reactants and catalyst are in the same phase.	In this catalysis, reactants and catalyst are in different phase.
2	In this, the system remains homogeneous during the reaction.	In this, the system remains heterogeneous during the reaction.
3	Mostly liquids and gases are used as homogeneous catalysts.	Mostly solid are used as heterogeneous catalysts.
4	Example: $2SO_{2(g)} + O_{2(g)} \xrightarrow{NO} 2SO_{3(g)}$ NO is a homogeneous catalyst.	Example: $2SO_{2(g)} + O_{2(g)} \xrightarrow{V_2O_5} 2SO_{3(g)}$ V_2O_5 is heterogeneous catalyst.

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Joan I. Senyk, Darrell D. Ebbing



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