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# **Os Segredos da Maleta Vermelha**

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**Scott C. Dulebohn**



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web the four different types of orbitals s p d and f have different shapes and one orbital can hold a maximum of two electrons the p d and f orbitals have different sublevels thus can hold more electrons as stated the electron configuration of each element is unique to its position on the periodic table

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web the electronic configuration of an atom in the quantum mechanical model is stated by listing the occupied orbitals in order of filling with the number of electrons in each orbital indicated by superscript in this notation the electronic configuration of sodium would be  $1s^2 2s^2 2p^6 3s^1$  distributed in the orbitals as

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web the orbital diagram of helium is  $1s^2$  the spin of an electron is specified by a fourth quantum number called the the two possible values of  $m_s$  are spin up and spin down the three principles that are needed to specify orbital occupancy for the electrons are pauli exclusion principle no orbital may have more than two electrons and they must have opposite spins

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web the electron configuration and the orbital diagram are following hydrogen is the noble gas helium which has an atomic number of 2 the helium atom contains two protons and two electrons the first electron has the same four quantum numbers as the hydrogen atom electron  $n=1, l=0, m_l=0, m_s=1/2$

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web chemistry questions and answers a write the electron configuration draw the orbital diagram determine the distinguishing electron and determine the 4 quantum numbers for the distinguishing electron of the element magnesium mg write electron configurations as  $1s^2 2s^2 2p^6 3s^2 3p^6$  with a space between each entry and no superscripts or rn

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web the easiest way to create electron configurations is using an electron configuration table which is a way of writing down the various orbitals available to electrons this table is easy to remember and it makes it possible to generate the electron configuration table for

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web this problem has been solved you ll get a detailed solution from a subject matter expert that helps you learn core concepts question write the complete electron configuration and orbital diagram for each of the following enter up to indicate an upwards pointing arrow down to indicate a downwards pointing arrow up down to indicate two

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web determine element from electron configuration determine number of electrons in an element s orbital s element groups to use select at least one s block alkali metals and alkaline earth metals p block nonmetals metalloids and other metals d

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web chemistry chemistry questions and answers part awrite the full electron configuration for the following  $Si^3+$  part bwrite the noble gas abbreviated electron configuration for the following  $Nb^4+$  part cdraw the orbital diagram for the following excited  $Ca^2+$  determine the number of unpaired electrons from your previous answer