Section 5.2 Review The Structure of the Atom

Multiple Choice

For each auestion below, select the letter of the Protect appropriate.

- 1. Which of the following statements about protons, neutrons, and electrons is traw?
 - (a) Protons and electrons have about the same
 - (b) Neutrons and electrons have equal but opposite charges.
 - (c) The mass of a proton is nearly 2000 times greater than the mass of an electron.
 - (d) The mass of an proton is nearly 2000 times greater than the mass of a neutron.
 - (e) Neutrons and electrons are found in the nucleus of an atom.
- 2. What is the atomic number of mercury, Hg?
 - 643 24 ·
 - dbn 800
 - (c) 120
 - (d) 200
 - (c) 200.6
- 3. What is the mass number of an atom of chlorine with 18 neutrons?
 - (a) I-
 - (b) 17.
 - 6cb 18
 - (d) 35
 - (c) 35.5
- 4. What element is represented by the Bohr-Rutherford diagram below?
 - (a) potassium, K.
 - (b) enleium, Ca-
 - (c) hydrogen, H
 - (d) argon, Ar.
 - (c) yttrium, Y

- 5. Which of the following statements about the massesof attoms is folse?
 - (a) Most of the mass of an atom is located in the
 - (b) Electrons contribute very little to the mass of ain attorn.
 - (c) Atoms of different elements have different PERSONAL.
 - (d) Most of the mass of an atom's nucleus is due to the mass of neutrons.
 - (e) The mass number of an atom is equal to the number of protons plus the number of neutrons.
- 6. A neutral atom has eight electrons and eight neutrons. Which of the following options correctly represents the atom?
 - (a) carbon-16
 - (b) oxygen-16
 - (c) oxygen-24
 - (d) oxygen-8
 - (c) sulfur-8
- 7. Which of the following symbols represents the element with an atomic number of 147
 - (a) C
 - Obb IN
 - felt Nil
 - (40) Si
 - tien S.

the statements is false? heutrons an atom has defines

- number delectrons an atom has defines
- The what element it is.

 (d) First the above are true.

 Cone of the above are true. ber of protons an atom has defines

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