

Potentiometry

Choose the correct answer:

| No | Sentence | answer |
|----|--------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------|--------|
| 1 | All of the following are related to reference electrode except: a) of known and constant potential b) obeys Nernst equation. c) of high resistance to permit current of electric current. d) insensitive to activity of analyte solution. | c |
| 2 | Potential of standard hydrogen electrode is: a) Zero b) related to H ⁺ concentration c) 0.059 PH d) A and B e) All of the above | d |
| 3 | The increase in chloride ion concentration in Ag/AgCl electrode leads to: a) increase the potential b) decrease the potential c) no effect d) none of them | b |
| 4 | Potential of calomel electrode is proportional to chloride ion concentration. a) directly b) inversely c) equal d) no relation | b |
| 5 | In silver/silver chloride electrode, as KCl concentration increases: a) silver ions increases, potential increases. b) silver ions decreases, potential increases. c) silver ions decreases, potential remains the same. d) silver ions decreases, potential decreases. | d |
| 6 | In calomel electrode, as KCl concentration increases: a) mercury(I) ions increases, potential increases. b) mercury(I) ions decreases, potential increases. c) mercury(I) ions decreases, potential remains the same. d) mercury(I) ions decreases, potential decreases. | d |
| 7 | What is the electrode potential for a half-cell consisting of a cadmium electrode immersed in a solution that is 0.0150 M in Cd ²⁺ ($E^\circ = -0.403$ V) a) 0.457 b) 0.403 c) -0.457 d) -0.201 | c |

Mcq On Potentiometry

Camilla Rothe

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