

## Assessment and Evaluation

ThoughtLab/ ExpressLab/ Investigation	Curriculum Expectations	Assessment Tools/Techniques	Achievement Chart Category	Learning Skills
ExpressLab: A Penny for Your Isotopes, p. 168	Understanding Basic Concepts <ul style="list-style-type: none"> <li>[OCR 1.62] explain the relationship between isotopic abundance and relative atomic mass</li> </ul>	<ul style="list-style-type: none"> <li>Assessment Checklist 5: Learning Skills (see "Assessment and Evaluation" in the front matter of <i>Teacher's Resource CD-ROM</i>)</li> </ul>	<ul style="list-style-type: none"> <li>Knowledge/ Understanding</li> </ul>	<ul style="list-style-type: none"> <li>Teamwork</li> </ul>

## Solutions to Practice Problems

Student Textbook page 170

See *Solutions Manual* for solutions to Practice Problems.

## Section Review Answers

Student Textbook page 170

- The average atomic mass is a weighted average of the atomic masses of the isotopes of potassium. Any given atom would have the atomic mass of one of the isotopes, not the average mass of all the isotopes.

2.

Isotope	Atomic Mass (u)	Percentage Abundance
Mg-24	23.985	78.78
Mg-25	24.985	10.13
Mg-26	25.983	11.17

$$\text{average atomic mass} = \frac{[78.78(23.985) + 10.13(24.985) + 11.17(25.983)] \text{ u}}{100}$$

$$\text{average atomic mass} = \frac{[1887.6 + 253.1 + 290.2] \text{ u}}{100}$$

$$\text{average atomic mass} = 24.31 \text{ u}$$

Therefore, the average atomic mass of magnesium is 24.31 u.

- It is not possible to solve the problem. There would be two unknowns in just one equation. Percentage abundance for one isotope is needed.
- Radioisotopes occur in nature along with the stable isotopes of an element. The nucleus of a radioisotope is continually altered due to the loss of alpha or beta particles. This in turn affects the atomic mass of that isotope as it undergoes radioactive decay. Changing masses such as this should not be taken into consideration when calculating the average atomic mass of an element. Another reason these isotopes are not used in calculating the average atomic mass of silver is that their natural abundance is usually so small that their contribution to the average atomic mass is insignificant.

# Mhr Chemistry Review Solutions Unit 5

**Christian G. Meyer**



## Mhr Chemistry Review Solutions Unit 5:

**Energy Research Abstracts** ,1981      **Scientific and Technical Aerospace Reports** ,1974 Lists citations with abstracts for aerospace related reports obtained from world wide sources and announces documents that have recently been entered into the NASA Scientific and Technical Information Database      Journal of Applied Chemistry ,1960      **Index to Scientific Reviews** ,1982      **Proceedings of the National Academy of Sciences of the United States of America** National Academy of Sciences (U.S.),2007      **Science Citation Index** ,1994 Vols for 1964 have guides and journal lists

**Physics Briefs** ,1980      **Government Reports Announcements & Index** ,1995      Pearson Chemistry Unit 5 Staley Wilbraham (Matta, Waterman),2022      **Study Guide and Solutions Manual to Accompany Organic Chemistry** G. Marc Loudon,Joseph G. Stowell,2003 This student Study Guide Solutions Manual acclaimed as one of the best in the field supplies not only answers but also detailed solutions to all text problems in Organic Chemistry Fourth Edition by G Marc Loudon Its Study Guide Links show students how to solve problems provide shortcuts to mastering particular topics and offer detailed discussions of concepts that students often find difficult Full chapter outlines a glossary of terms and reaction reviews are provided      *Chemistry* Nicholas Drapela,Alison Hyslop,2005      *Practice Makes Perfect Chemistry Review and Workbook, Second Edition* Marian DeWane,Heather Hattori,2018-12-28 The Winning Equation for Success in Chemistry is Practice Practice Practice This book will help you apply concepts and see how chemistry topics are interconnected Inside are numerous lessons to help you better understand the subject These lessons are accompanied by dozens of exercises to practice what you ve learned along with a complete answer key to check your work Throughout this book you will learn the terms to help you understand chemistry and you will expand your knowledge of the subject through hundreds of sample questions and their solutions With the lessons in this book you will find it easier than ever to grasp chemistry concepts And with a variety of exercises for practice you will gain confidence using your growing chemistry skills in your classwork and on exams YOU LL BE ON YOUR WAY TO MASTERING THESE TOPICS AND MORE Atomic structure The periodic table Chemical formulas Chemical reactions Mass and mole relationships Gas laws Solutions Acids and bases Thermochemistry A brand new chapter on the structure of molecules      Math Review ToolKit for General Chemistry Gary Leighton Long,Ralph H. Petrucci,Sharon D. Long,F. Geoffrey Herring,Jeffrey D. Madura,Carey Bissonnette,2010-04-12      **Study Guide and Solutions Manual for McMurry's Organic Chemistry, Fifth Edition** Susan McMurry,2000 Provides answers and explanations to all in text and end of chapter exercises Also includes summaries of name reactions functional group synthesis and reactions lists of reagents and abbreviations and articles on topics ranging from infrared absorption frequencies to the Nobel Price winners in Chemistry This edition now includes all new artwork expanded in text problems summary quizzes approximately every three chapters more detailed explanations in solutions and chapter outlines      Cracking the Virginia SOL Paul Foglino,Princeton Review (Firm),2000 The Princeton Review realizes that acing the Chemistry exam is very

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