

Chapter-01 - (Page 19 of 21)

beeducated

(a) N₂ (b) O₂ (c) Cl₂ (d) S

10. Which gas is used to manufacture king of chemicals sulphuric acid?

(a) N₂ (c) Cl₂ (b) O₂ (d) S

11. Equilibrium constant has no unit when number of moles of reactants and products are:

(a) same

(b) different

(c) both a & b

(d) none

 For reactions having large Kc value, the reaction proceeds to;

- (a) completion
- (b) equilibrium state
- (c) back ward
- (d) None

13. The characteristics of reversible reactions are the following except:

- (a) products never recombine to form reactants
- (b) they never complete
- (c) they proceed in both ways
- (d) they have a double arrow between reactants and products

14. In the lime kiln, the reaction

 $CaCO_{3(s)} \longrightarrow CaO_{(s)} + CO_{2(g)}$ goes to

completion because;

- (a) of high temperature
- (b) CaO is more stable than CaCO₃
- (c) CO₂ escapes continuously
- (d) CaO is not dissociated

15. For the reaction,

 $2A_{(g)} + B_{(g)} \Longrightarrow 3C_{(g)}$ the expression

for the equilibrium constant is:

(a) $\frac{[2A][B]}{[3C]}$ (b) $\frac{[A]^2[B]}{[C]^3}$

(c) $\frac{[3C]}{[2A][B]}$ (d) $\frac{[C]^3}{[A]^2[B]}$

16. When a system is at equilibrium states?

- (a) the concentration of reactants and products becomes equal
- (b) the opposing reactions (forward and reverse) stop
- (c) the rate of the reverse reaction becomes very low
- (d) the rates of the forward and reverse reactions becomes equal.

17. Which one of the following statements is not correct about active mass?

- (a) rate of reaction is directly proportional to active mass.
- (b) active mass is taken in molar concentration
- (c) active mass is represented by square brackets
- (d) active mass means total mass of substances

18. When the magnitude of K_c is very large it indicates;

- (a) reaction mixture consists of almost all products
- (b) reaction mixture has almost all reactants
- (c) reaction has not gone to completion
- (d) reaction mixture has negligible products

 When the magnitude of K_c is very small it indicates;

(a) equilibrium will never establish

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