

Thermochemistry Chemistry Questions with Solutions

Q-1: Given that standard molar enthalpies of formation of NO(g) and NO₂(g) are respectively 90.3 kJ/mol and 33.2 kJ/mol. Calculate the enthalpy change for the reaction $2NO(g) + O_2(g) \rightarrow 2NO_2(g)$.

Answer: ΔH = ΣΔH_{protes}- ΣΔH_{macters} ΔH = 2(33.2)-[2(90.3)+0] ΔH = -114.2kJ

Note: Enthalpy of formation of free elements like O₂ is always 0.

Q-2: The enthalpy of formation of AgCI is obtained from the enthalpy change from which one of the following process?

- a) Ag*(aq) +Cr(aq) →AgCl(s)
- b) Ag(s) + ½ Cl₂(g) →AgCl(s)
- c) AgCl(s) →Ag(s) + ½ Cl₂(g)
- d) Ag(s) +AuCl →Au(s) + AgCl(s)

Answer: b) $Ag(s) + \frac{1}{2} Cl_2(g) \rightarrow AgCl(s)$

<u>Explanation</u>: The enthalpy change accompanying the formation of one mole of a pure substance in its standard state from the constituents elements in their standard state is called enthalpy of formation. For Ag, the standard state is solid and for chlorine, it is gas. Thus, the process that gives the enthalpy of formation of AgCl is: $Ag(s) + \frac{1}{2} Cl_2(g) \rightarrow AgCl(s)$

Q-3: ∆H and ∆E for the reaction, Fe₂O₃ (s) +3H₂(g) → 2Fe(s) + H₂O(l) at constant temperature are related as

- a) ΔH = ΔE
- b) $\Delta H = \Delta E + RT$
- c) $\Delta H = \Delta E + 3RT$
- d) $\Delta H = \Delta E 3RT$

Answer: d) $\Delta H = \Delta E - 3RT$

Explanation: For any chemical reaction,

 $\Delta H = \Delta E + \Delta_{to}RT$

Where, Δ_{eg} = total number of moles of gaseous product- total number of moles of gaseous reactants.

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